

13.1**THE NATURE OF GASES****Section Review****Objectives**

- Describe the assumptions of the kinetic theory as it applies to gases
- Interpret gas pressure in terms of kinetic theory
- Define the relationship between Kelvin temperature and average kinetic energy

Vocabulary

- | | | |
|------------------|------------------------|-----------------------------|
| • kinetic energy | • vacuum | • pascal (Pa) |
| • kinetic theory | • atmospheric pressure | • standard atmosphere (atm) |
| • gas pressure | • barometer | |

Part A Completion

Use this completion exercise to check your understanding of the concepts and terms that are introduced in this section. Each blank can be completed with a term, short phrase, or number.

The kinetic theory describes the _____ of particles in matter and the forces of attraction between them. The theory assumes that the volume occupied by a gas is mostly _____, that the particles of gas are relatively _____, move _____ of each other, and are in constant _____ motion. The _____ between particles are perfectly elastic so that the total _____ remains constant. Gas pressure results from the simultaneous collisions of billions of particles with an object. Barometers are used to measure _____ pressure. Standard conditions are defined as a temperature of _____ and a pressure of _____.

1. _____
2. _____
3. _____
4. _____
5. _____
6. _____
7. _____
8. _____
9. _____
10. _____

8
9

Part B True-False

Classify each of these statements as always true, AT; sometimes true, ST; or never true, NT.

- _____ 11. Atmospheric pressure is 760 mm Hg.
- _____ 12. The SI unit of pressure is the pascal.

Modern Chemistry Gases Review Section 2 Answers

Yijin Wang



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