

balance nuclear reaction equations

aluminum-26 undergoes positron emission
potassium-40 undergoes beta decay (electron)
uranium-238 undergoes alpha decay



electron configurations

Write the electron configuration for C.

Write the electron configuration for Cr.

Write the shorthand notation for Bi.



NOMENCLATURE [naming compounds, writing formulas]

Write the formula for each of the following compounds:

- a. Sulfur hexafluoride SF_6
- b. Lithium nitride Li_3N
- c. Chromium (III) carbonate $\text{Cr}_2(\text{CO}_3)_3$
- d. Tin (II) chloride SnCl_2
- e. Ammonium acetate $\text{NH}_4\text{C}_2\text{H}_3\text{O}_2$ (or $\text{CH}_3\text{COONH}_4$)
- f. Mercury (I) chloride Hg_2Cl_2
- g. Potassium bromate KBrO_3
- h. Hydrosulfuric acid $\text{H}_2\text{S(aq)}$
- i. Chloric acid HClO_3
- j. Sulfurous acid H_2SO_3

Name each of the following compounds.

- a. CuSO_4 Copper (II) sulfate
- b. AlF_3 aluminum fluoride
- c. HI hydrogen iodide
- d. NO nitrogen monoxide
- e. $\text{H}_2\text{Se(aq)}$ hydroselenic acid
- f. HNO_3 nitric acid
- g. HNO_2 nitrous acid
- h. NaHSO_4 sodium hydrogen sulfate

Half-life

Tritium (H-3) is a radioactive isotope of hydrogen with a half-life of 12.3 years. How long would it take for a 40.0 g sample to decay down to 1.25 g?

$$40 \xrightarrow{1} 20 \xrightarrow{2} 10 \xrightarrow{3} 5 \xrightarrow{4} 2.5 \xrightarrow{5} 1.25$$

Fe-61 has a half-life of 6.00 min. Of a 100.0 mg sample, how much will remain after 18.0 min?

$$3 \text{ half lives} \quad 100 \xrightarrow{1} 50 \xrightarrow{2} 25 \xrightarrow{3} 12.5 \text{ mg}$$

Average atomic mass

The element copper has naturally occurring isotopes with mass numbers of 63 and 65.

The relative abundance and atomic masses are 69.2% for a mass of 63amu and 30.8% for a mass of 65amu. Calculate the average atomic mass of copper.

$$\begin{array}{rcl} 63 \times .692 & = & 43.596 \rightarrow 44 \\ 65 \times .308 & = & 20.02 \rightarrow 20. \\ \hline & + & \\ & 63.62 & \rightarrow 64 \end{array}$$

Modern Chemistry Chapter 13 2 Review Answers

Norma K Dunlap, Donna M Huryn



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Modern Chemistry Holt Rinehart & Winston, Holt, Rinehart and Winston Staff, 2001 **Medicinal Chemistry** Norma K Dunlap, Donna M Huryn, 2025-11-28 This book provides a concise introduction to the medicinal chemistry of therapeutically active compounds. Written in an accessible style, it offers a brief overview of the drug development process from discovery to approved drug. Emphasizing medicinal chemistry strategies, the content is organized by target as well as therapeutic area, drug classes. It covers existing drugs and the discovery process that led to their development. Aimed at students and professionals interested in drug discovery, Medicinal Chemistry requires little prior knowledge, making it ideal for classroom and training seminar use. Written by experts in the field, this outstanding new edition explains medicinal chemistry on chemical mechanism terms, allowing for a more interesting and user-friendly text. Includes valuable insights toward the various pathways taken by the pharmaceutical industry in drug discoveries. Enhanced text includes questions raised and suggestions made from students in the author's medicinal chemistry class. Content provided to enhance the text includes a solutions manual, PowerPoint slides for lecture outlines, and especially for the earlier chapters, a selection of videos. An accompanying website will be updated to include additional material: 1. Journal Club exercises, current journal articles with guiding questions for student analysis; two per chapter; 2. additional case studies; 3. organic synthesis analysis, selected syntheses, both discovery and process, will be discussed; and 4. updates on current advances in the field of medicinal chemistry. The book benefits a diverse audience, including upper-level undergraduates and graduates who are studying in the fields of medicinal chemistry and drug discovery, as well as scientists working in the pharmaceutical industry. **Independent and Weekly Review**, 1914 *Modern Enolate Chemistry* Manfred Braun, 2016-02-23 Authored by one of the world's leading synthetic chemists in the field, this reference presents modern enolate chemistry with an emphasis on metal O-enolates in asymmetric synthesis. While great care is taken to cover novel, successful concepts, such as classical methods, as the famous Evans enolates are equally highlighted. Throughout the book, representative reaction procedures are presented, thus helping readers to find the best solution for their own synthetic problem. Of high interest to synthetic chemists in academia, as well as the pharmaceuticals, agrochemicals, and fine chemicals industries. **Foundations of College Chemistry, Student Solutions Manual** Morris Hein, Susan Arena, Cary Willard, 2022-03-09 Check your work and better retain the lessons of Foundations of College Chemistry. The Foundations of College Chemistry 16th Edition Student Solutions Manual is an essential study resource and aid that allows users of the widely read text to check their work against the detailed solutions contained in this book. This Solutions Manual unlocks the full value of the practice problem and exercise-focused 16th edition of Foundations of College Chemistry. Students will be able to study efficiently and effectively and learn from common errors with the additional guidance offered by this book. [Solutions Guide for Introductory Chemistry](#) James F. Hall, 1996

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