

SECTION-1 (SINGLE ANSWER TYPE QUESTIONS)

31. Number of atoms in 560 g of Fe (atomic mass 56 g mol^{-1}) is
 i) twice that of 70 g N
 ii) half that of 20 g H
 iii) half that of 40 g He
 a) i, ii b) ii, iii c) i, ii, iii d) only iii
32. How many moles of magnesium phosphate, $\text{Mg}_3(\text{PO}_4)_2$ will contain 2.5 mol of oxygen atoms?
 a) 3.2 b) 3.125×10^{-1} c) 3.125×10^{-2} d) 0.32
33. Experimentally it was found that a metal oxide has formula $\text{M}_{0.95}\text{O}$. Metal M, is present as M^{2+} and M^{3+} in its oxide. Fraction of the metal which exists as M^{3+} would be
 a) 4.08 % b) 6.05 % c) 95.92 % d) 93.95 %
34. The percentage composition of carbon by mole in methane is
 a) 75 % b) 80 % c) 25 % d) 20 %
35. A 5.2 molal aqueous solution of methyl alcohol CH_3OH is supplied. What is the mole fraction of methyl alcohol in the solution?
 a) 0.05 b) 0.090 c) 0.052 d) 0.2
36. A solution of sodium sulfate contains 92 g of Na^+ ions per kilogram of water. The molality of Na^+ ions in that solution in mol kg^{-1} is
 a) 12 b) 4 c) 8 d) 16
37. 8 g of NaOH is dissolved in 18 g of H_2O . Mole fraction of NaOH in solution and molality (in mol kg^{-1}) of the solution respectively are
 a) 0.2, 22.20 b) 0.2, 11.11 c) 0.167, 11.11 d) 0.167, 22.20
38. The amount of oxygen in 3.6 mol of water is
 a) 64.8 g b) 57.6 g c) 28.8 g d) 18.4 g

39. 25 mL of a solution of barium hydroxide on titration with 0.1 molar solution of hydrochloric acid gave a titre value of 35 mL. The molarity of barium hydroxide solution was
 a) 0.07 b) 0.14 c) 0.28 d) 0.35
40. The volume of 0.1 N dibasic acid sufficient to neutralize 1 g of a base that furnishes $0.04 \text{ mol}^{-1} \text{OH}^-$ in aqueous solution is
 a) 400 mL b) 600 mL c) 200 mL d) 800 mL
41. A gaseous compound of nitrogen and hydrogen contains 12.5% (by mass) of hydrogen. The density of the compound relative to hydrogen is 16. The molecular formula of the compound is
 a) NH_2 b) N_3H c) NH_3 d) N_2H_4
42. The amount of sugar ($\text{C}_{12}\text{H}_{22}\text{O}_{11}$) required to prepare 2 L of its 0.1 M aqueous solution is
 a) 136.8 g b) 17.1 g c) 68.4 g d) 34.2 g
43. The molarity of a solution obtained by mixing 750 mL of 0.5 (M) HCl with 250 mL of 2 (M) HCl will be
 a) 1.00 M b) 1.75 M c) 0.975 M d) 0.875 M
44. For a reaction, $\text{N}_2(\text{g}) + 3\text{H}_2(\text{g}) \rightarrow 2\text{NH}_3(\text{g})$
 Identify dihydrogen (H_2) as a limiting reagent in the following reaction mixtures.
 a) 56 g of N_2 + 10 g of H_2 b) 35 g of N_2 + 8 g of H_2
 c) 28 g of N_2 + 6 g of H_2 d) 14 g of N_2 + 4 g of H_2
45. For the following reaction, the mass of water produced from 445 g of $\text{C}_{27}\text{H}_{110}\text{O}_6$ is
 $2\text{C}_{27}\text{H}_{110}\text{O}_6(\text{s}) + 163\text{O}_2(\text{g}) \rightarrow 114\text{CO}_2(\text{g}) + 110\text{H}_2\text{O}(\text{l})$
 a) 490 g b) 485 g c) 495 g d) 480 g
46. In an organic compound of molar mass 108 g mol^{-1} C, H and N atoms are present in 9: 1: 3.5 by weight. Molecular formula can be

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Table of Contents Modern Chemistry Section 1 Question And Answer

1. Understanding the eBook Modern Chemistry Section 1 Question And Answer
 - The Rise of Digital Reading Modern Chemistry Section 1 Question And Answer
 - Advantages of eBooks Over Traditional Books
2. Identifying Modern Chemistry Section 1 Question And Answer
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Modern Chemistry Section 1 Question And Answer
 - User-Friendly Interface
4. Exploring eBook Recommendations from Modern Chemistry Section 1 Question And Answer
 - Personalized Recommendations
 - Modern Chemistry Section 1 Question And Answer User Reviews and Ratings
 - Modern Chemistry Section 1 Question And Answer and Bestseller Lists
5. Accessing Modern Chemistry Section 1 Question And Answer Free and Paid eBooks

- Modern Chemistry Section 1 Question And Answer Public Domain eBooks
 - Modern Chemistry Section 1 Question And Answer eBook Subscription Services
 - Modern Chemistry Section 1 Question And Answer Budget-Friendly Options
6. Navigating Modern Chemistry Section 1 Question And Answer eBook Formats
- ePub, PDF, MOBI, and More
 - Modern Chemistry Section 1 Question And Answer Compatibility with Devices
 - Modern Chemistry Section 1 Question And Answer Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Modern Chemistry Section 1 Question And Answer
 - Highlighting and Note-Taking Modern Chemistry Section 1 Question And Answer
 - Interactive Elements Modern Chemistry Section 1 Question And Answer
8. Staying Engaged with Modern Chemistry Section 1 Question And Answer
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Modern Chemistry Section 1 Question And Answer
9. Balancing eBooks and Physical Books Modern Chemistry Section 1 Question And Answer
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Modern Chemistry Section 1 Question And Answer
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Modern Chemistry Section 1 Question And Answer
- Setting Reading Goals Modern Chemistry Section 1 Question And Answer
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Modern Chemistry Section 1 Question And Answer
- Fact-Checking eBook Content of Modern Chemistry Section 1 Question And Answer
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development

- Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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