

## Stoichiometry Worksheet # 1

1) Glucose is used as a source of energy by the human body. The overall reaction in the body is



Calculate the number of moles of oxygen needed to oxidize 12.5 g of  $\text{C}_6\text{H}_{12}\text{O}_6$ .

2) Ammonia is synthesized from hydrogen and nitrogen according to the following equation.



If an excess of nitrogen is reacted with 3.41 moles of hydrogen gas, how many grams of ammonia ( $\text{NH}_3$ ) can be produced?

3) Assume that in the decomposition of potassium chlorate,  $\text{KClO}_3$ ,



80.5 moles of  $\text{O}_2$  form. How many grams of potassium chlorite, the other product would be formed?

4) In a single displacement reaction,



9.23 moles of aluminum react with excess hydrochloric acid. How many moles of hydrogen will be produced?

5) The compound "Cisplatin"  $\text{PtCl}_2(\text{NH}_3)_2$ , has been found to be effective in treating some types of cancer. It can be synthesized using the following reaction



A. How many moles of "cisplatin" can be produced from 2.50 g  $\text{K}_2\text{PtCl}_4$ ?

B. How many moles of  $\text{NH}_3$  would be needed if 3.8 moles of  $\text{K}_2\text{PtCl}_4$  were present?

# Modern Chemistry Stoichiometry Review Answer Key

**G Thomas**



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## **Table of Contents Modern Chemistry Stoichiometry Review Answer Key**

1. Understanding the eBook Modern Chemistry Stoichiometry Review Answer Key
  - The Rise of Digital Reading Modern Chemistry Stoichiometry Review Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Modern Chemistry Stoichiometry Review Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Modern Chemistry Stoichiometry Review Answer Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Modern Chemistry Stoichiometry Review Answer Key
  - Personalized Recommendations
  - Modern Chemistry Stoichiometry Review Answer Key User Reviews and Ratings
  - Modern Chemistry Stoichiometry Review Answer Key and Bestseller Lists
5. Accessing Modern Chemistry Stoichiometry Review Answer Key Free and Paid eBooks
  - Modern Chemistry Stoichiometry Review Answer Key Public Domain eBooks
  - Modern Chemistry Stoichiometry Review Answer Key eBook Subscription Services
  - Modern Chemistry Stoichiometry Review Answer Key Budget-Friendly Options

6. Navigating Modern Chemistry Stoichiometry Review Answer Key eBook Formats
  - ePub, PDF, MOBI, and More
  - Modern Chemistry Stoichiometry Review Answer Key Compatibility with Devices
  - Modern Chemistry Stoichiometry Review Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Modern Chemistry Stoichiometry Review Answer Key
  - Highlighting and Note-Taking Modern Chemistry Stoichiometry Review Answer Key
  - Interactive Elements Modern Chemistry Stoichiometry Review Answer Key
8. Staying Engaged with Modern Chemistry Stoichiometry Review Answer Key
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Modern Chemistry Stoichiometry Review Answer Key
9. Balancing eBooks and Physical Books Modern Chemistry Stoichiometry Review Answer Key
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Modern Chemistry Stoichiometry Review Answer Key
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Modern Chemistry Stoichiometry Review Answer Key
  - Setting Reading Goals Modern Chemistry Stoichiometry Review Answer Key
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Modern Chemistry Stoichiometry Review Answer Key
  - Fact-Checking eBook Content of Modern Chemistry Stoichiometry Review Answer Key
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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