

Chemical Quantities Outline

(The Molar Mass)

Molar Mass

- Atomic Mass is the mass of a single atom expressed in amu; relative values based on the isotopes; (decimals) on the periodic table
- Molar mass is the mass of a mole of an element/compound/molecule; equal to the atomic mass
Units are g/mole (grams per mole)
- Example:

Carbon's atomic mass 12.01 amu and its molar mass is 12.01 g/mol

How to find the mass of a mole of a compound

- You must know the chemical formula
- Then add the atomic masses of the atoms that make up that compound
- Example: Sulfur Trioxide (SO_3)

1 Sulfur $\rightarrow 32.06 \text{ g (1)} = 32.06 \text{ g/mol S}$
 3 Oxygens $\rightarrow 16.00 \text{ g (3)} = 48.00 \text{ g/mol O}$
 $\rightarrow 80.06 \text{ g/mol SO}_3$

The decomposition of hydrogen peroxide (H_2O_2) provides sufficient energy to launch a rocket. What is the molar mass of hydrogen peroxide?

2 Hydrogen
2 Oxygen

$2(1.008 \text{ g/mol}) \text{ H} = 2.016 \text{ g/mol H}$
 $2(16.00 \text{ g/mol}) \text{ O} = 32.00 \text{ g/mol O}$
 $= 34.02 \text{ g/mol H}_2\text{O}_2$

Find the molar mass of calcium nitrate

1 Calcium
2 Nitrogen
6 Oxygen
 $\text{Ca}(\text{NO}_3)_2$
 $1(40.08 \text{ g/mol Ca}) + 2(14.01 \text{ g/mol N}) + 6(16.00 \text{ g/mol O})$
 $40.08 \text{ g/mol} + 28.02 \text{ g/mol} + 96 \text{ g/mol} = 164.1 \text{ g/mol Ca}(\text{NO}_3)_2$

Mole to Mass Relationship

- Molar mass of a substance is the mass in grams per mole of that substance
- Use the molar mass of an element or compound to convert between mass and the mole
- Conversion factors: $\text{molar mass} = \frac{\text{g}}{\text{mol}}$ or $\frac{\text{mol}}{\text{g}}$

Molar Mass And Chemical Quantities Answer Key

F Rizvi



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