

## Molarity and Dilutions Practice - HN

## Part 1: Molarity

$$M = \frac{\text{moles of solute}}{\text{Volume of Solution (L)}} = \frac{\text{Mols}}{\text{L}}$$

1. What is the molarity of a 0.30 liter solution containing 0.50 moles of sodium chloride, NaCl?

$$M = \frac{.5 \text{ mol}}{.3 \text{ L}}$$

$$M = \boxed{1.67 \text{ M}}$$

2. Calculate the molarity of 0.289 moles of Iron (III) Chloride, FeCl<sub>3</sub>, dissolved in 120 mL of solution.

$$\textcircled{1} \quad 120 \text{ mL} \times \frac{1 \text{ L}}{1000 \text{ mL}} = .12 \text{ L}$$

mL → L

$$\textcircled{2} \quad M = \frac{.289 \text{ mol}}{.12 \text{ L}} = \boxed{2.41 \text{ M}}$$

3. What is the molarity of 0.5 grams of sodium chloride, NaCl, dissolved to make 50 mL of solution?

$$\textcircled{1} \quad .5 \text{ g} \times \frac{1 \text{ mol}}{58.44 \text{ g}} = .009 \text{ mol}$$

g → mol

$$\textcircled{2} \quad 50 \text{ mL} \times \frac{1 \text{ L}}{1000 \text{ mL}} = .05 \text{ L}$$

$$\textcircled{3} \quad M = \frac{.009 \text{ mol}}{.05 \text{ L}} = \boxed{.18 \text{ M}}$$

4. Calculate the molarity of 734 grams of lithium sulfate, Li<sub>2</sub>SO<sub>4</sub>, dissolved in 2,500 mL of solution.

$$\textcircled{1} \quad 734 \text{ g} \times \frac{1 \text{ mol}}{109.95 \text{ g}} = 6.68 \text{ mol}$$

$$\textcircled{2} \quad 2500 \text{ mL} \times \frac{1 \text{ L}}{1000 \text{ mL}} = 2.5 \text{ L}$$

$$\textcircled{3} \quad M = \frac{6.68 \text{ mol}}{2.5 \text{ L}} = \boxed{2.67 \text{ M}}$$

5. How many grams of Ca(OH)<sub>2</sub> are needed to produce 500 mL of a 1.66M Ca(OH)<sub>2</sub> solution?

$$\textcircled{1} \quad 500 \text{ mL} \times \frac{1 \text{ L}}{1000 \text{ mL}} = .5 \text{ L}$$

$$\textcircled{2} \quad 1.66 \text{ M} = \frac{\text{mol}}{.5 \text{ L}}$$

mol = .83 mol

$$\textcircled{3} \quad .83 \text{ mol} \times \frac{74.1 \text{ g}}{1 \text{ mol}} = \boxed{61.5 \text{ g Ca(OH)}_2}$$

6. How many grams of KOH are required to make 125 mL of a 0.50M solution?

$$\textcircled{1} \quad 125 \text{ mL} \times \frac{1 \text{ L}}{1000 \text{ mL}} = .125 \text{ L}$$

$$\textcircled{2} \quad .5 \text{ M} = \frac{\text{mol}}{.125}$$

mol = .063 mol

$$\textcircled{3} \quad .063 \text{ mol} \times \frac{56.108 \text{ g}}{1 \text{ mol}} = \boxed{3.53 \text{ g}}$$

7. How many grams of Na<sub>2</sub>SO<sub>4</sub> are required to make 575 mL of a 0.60M solution?

$$\textcircled{1} \quad 575 \text{ mL} \times \frac{1 \text{ L}}{1000 \text{ mL}} = .575 \text{ L}$$

$$\textcircled{2} \quad .6 \text{ M} = \frac{\text{mol}}{.575 \text{ L}}$$

mol = .345

$$\textcircled{3} \quad .345 \text{ mol} \times \frac{142.04 \text{ g}}{1 \text{ mol}} = \boxed{49.01 \text{ g}}$$

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