



Figure 1 The combustion of hydrogen fuel was used to launch the space shuttle into orbit. Aboard the shuttle, hydrogen fuel cells used the same chemical reaction to provide astronauts with electricity and water.

CAREER LINK

Chemical process engineers plan which reactants should be mixed, under which conditions, to give the optimum yield and produce the smallest quantity of unwanted by-products. To find out about this career,



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The reaction of hydrogen and oxygen has been an important energy source for the space program (Figure 1). Hydrogen combines with oxygen in a synthesis reaction to form water and a great deal of energy:



As you discovered in Investigation 4.4.1 Hydrogen Blast-Off in Unit 3, this reaction releases the most energy per mole of reactants when the volumes of hydrogen and oxygen are in a 2:1 ratio.

Ratios in Chemical Equations

A chemical equation is similar to a recipe, with the chemical formulas indicating the “ingredients” of the reaction. The coefficients in the equation give the ratio of one chemical to another chemical used in the reaction. However, these are ratios of numbers of entities rather than masses. For example, the equation



states that hydrogen molecules and oxygen molecules combine in a 2:1 ratio. This implies that 2 molecules of hydrogen are required to completely react with 1 molecule of oxygen. But what if more hydrogen is present? The amount of oxygen required to use up this extra hydrogen is determined using the 2:1 ratio. The amount of water produced is also determined by this ratio (Figure 2).

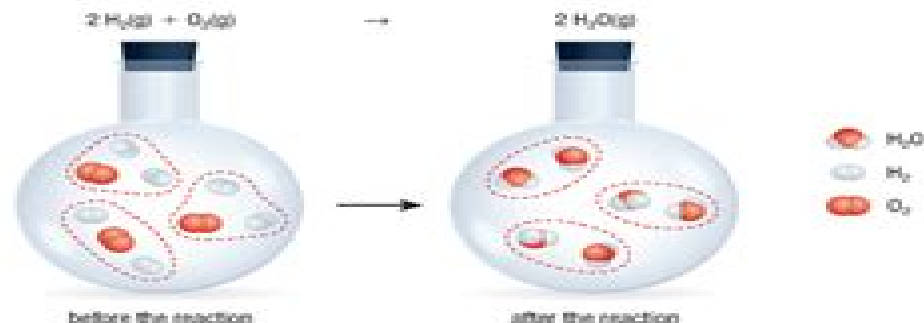


Figure 2 The synthesis of water from its elements. The coefficients in the chemical equation for the reaction give the ratio in which hydrogen and oxygen combine. In this case, 6 hydrogen molecules react with 3 oxygen molecules to form 6 water molecules.

Working with ratios is an essential skill in chemistry. The following analogy illustrates the reasoning used when working with ratios in chemical equations.

A Ratio Analogy

S'mores are sweet snacks that are fun to make over a campfire. They probably got their name from people asking, “Can I have some more?” Each one consists of two graham crackers sandwiching four chocolate chips and one marshmallow.

To make this recipe analogous to a chemical reaction, we will give each ingredient and product a chemical symbol:

- graham cracker, Gc
- chocolate chip, Cc
- marshmallow, Mm
- s'more, Sm

Mole Ratios Model 1 A Chemical Reaction Answers

K Morrison



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