

1/5/16

## Mole Ratio Worksheet

Key

① Given this equation:  $\text{N}_2 + 3\text{H}_2 \rightarrow 2\text{NH}_3$ , write the following molar ratios:

- a)  $\text{N}_2 / \text{H}_2$  1:3  
 b)  $\text{N}_2 / \text{NH}_3$  1:2  
 c)  $\text{H}_2 / \text{NH}_3$  3:2

② Given the following equation:  $8\text{H}_2 + \text{S}_8 \rightarrow 8\text{H}_2\text{S}$ , write the following molar ratios:

- a)  $\text{H}_2 / \text{H}_2\text{S}$  8:8  $\rightarrow$  1:1 (reduced)  
 b)  $\text{H}_2 / \text{S}_8$  8:1  
 c)  $\text{H}_2\text{S} / \text{S}_8$  8:1

③ Answer the following questions for this equation:  $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$

- a) What is the  $\text{H}_2 / \text{H}_2\text{O}$  molar ratio? 2:2  $\rightarrow$  1:1 (reduced)  
 b) Suppose you had 20 moles of  $\text{H}_2$  on hand and plenty of  $\text{O}_2$ , how many moles of  $\text{H}_2\text{O}$  could you make?  

$$\frac{20 \text{ mole H}_2}{2 \text{ H}_2} \times \frac{2 \text{ H}_2\text{O}}{2 \text{ H}_2} = \boxed{20 \text{ mole H}_2\text{O}}$$
  
 c) What is the  $\text{O}_2 / \text{H}_2\text{O}$  molar ratio?  
 1:2  
 d) Suppose you had 20 moles of  $\text{O}_2$  and enough  $\text{H}_2$ , how many moles of  $\text{H}_2\text{O}$  could you make?  

$$\frac{20 \text{ mol O}_2}{1 \text{ O}_2} \times \frac{2 \text{ H}_2\text{O}}{2 \text{ H}_2} = \boxed{40 \text{ mol H}_2\text{O}}$$

4) Use this equation:  $\text{N}_2 + 3\text{H}_2 \rightarrow 2\text{NH}_3$ , for the following problems

- a) If you used 1 mole of  $\text{N}_2$ , how many moles of  $\text{NH}_3$  could be produced?  $\boxed{2}$   
 b) If 10 moles of  $\text{NH}_3$  were produced, how many moles of  $\text{N}_2$  would be required?  

$$\frac{10 \text{ mol NH}_3}{2 \text{ NH}_3} \times \frac{1 \text{ N}_2}{1 \text{ N}_2} = \boxed{5 \text{ mol N}_2}$$
  
 c) If 3.00 moles of  $\text{H}_2$  were used, how many moles of  $\text{NH}_3$  would be made?  

$$\frac{3 \text{ mol H}_2}{3 \text{ H}_2} \times \frac{2 \text{ NH}_3}{2 \text{ NH}_3} = \boxed{2 \text{ mol NH}_3}$$
  
 d) If 0.600 moles of  $\text{NH}_3$  were produced, how many moles of  $\text{H}_2$  are required?  

$$\frac{0.6 \text{ mol NH}_3}{2 \text{ NH}_3} \times \frac{3 \text{ H}_2}{2 \text{ NH}_3} = \boxed{0.9 \text{ mol H}_2}$$

# Mole Ratio Answer Key

**Shasha Hu**

A red circular graphic with a gradient, appearing as a stylized arrow or a decorative element, located to the right of the author's name.

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**Teaching Science for Understanding** James Joseph Gallagher,2007 Offers middle and high school science teachers practical advice on how they can teach their students key concepts while building their understanding of the subject through various levels of learning activities

*Holt Chemistry* Ralph Thomas Myers,2004 **Study Guide, Chemical Principles, Fifth Edition, Zumdahl** Paul B. Kelter,Steven S. Zumdahl,2004-04

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    **Instructors Resource Manual** John McMurry, Robert A. Pribush, Robert Fay, 2000-12-27

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