

Unit 8 Worksheet 1: Mole relationships

For each of the problems below:

- Write the balanced chemical equation.
- Identify what is given (with units) and what you want to find (with units).
- Use coefficients from balanced equation to determine mole ratio.
- Show set up (organize it!).

- Hydrogen sulfide gas, which smells like rotten eggs, burns in air to produce sulfur dioxide and water. How many moles of oxygen gas would be needed to completely burn 8 moles of hydrogen sulfide?

Equation: $\underline{\quad 2 \quad} \text{H}_2\text{S}_{(g)} + \underline{\quad 3 \quad} \text{O}_{2(g)} \rightarrow \underline{\quad 2 \quad} \text{SO}_{2(g)} + \underline{\quad 2 \quad} \text{H}_2\text{O}_{(g)}$

Before $\underline{\quad 2 \quad}$ $\underline{\quad \text{ex} \quad}$ $\underline{\quad 0 \quad}$ $\underline{\quad 0 \quad}$ 3.0 mol O₂

Change $\underline{\quad -2 \quad}$ $\underline{\quad -3 \quad}$ $\underline{\quad 2 \quad}$ $\underline{\quad 2 \quad}$

After $\underline{\quad 0 \quad}$ $\underline{\quad \text{ex} \quad}$ $\underline{\quad 2 \quad}$ $\underline{\quad 2 \quad}$

- Propane, C₃H₈, burns in air to form carbon dioxide and water. If 12 moles of carbon dioxide are formed, how many moles of propane were burned?

Equation: $2\text{Al} + 6\text{HCl} \rightarrow 2\text{AlCl}_3 + 3\text{H}_2$

Before 0.5 $\underline{\quad \text{ex} \quad}$ 0 0 1.50mol HCL

Change -0.5 -1.5 +0.5 +0.5

After 0 $\underline{\quad \text{ex} \quad}$ 0.5 0.75

- Ammonia, NH₃, for fertilizer is made by causing hydrogen and nitrogen to react at high temperature and pressure. How many moles of ammonia can be made from 0.15 moles of nitrogen gas?

Equation: $3\text{H}_2 + \text{N}_2 \rightarrow 2\text{NH}_3$

Before $\underline{\quad \text{ex} \quad}$ 1.25 0 2.50 mole

Change -3.75 -1.25 +2.5

After $\underline{\quad \text{ex} \quad}$ 0 2.5

Mole Relationships Answers

Rosina Ehmann

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

Mole Relationships Answers:

Basic Concepts of Chemistry, 9e Study Guide and Solutions Manual Leo J. Malone, Theodore O. Dolter, 2012-01-03
The 9th edition of Malone's Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment. New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections, assessment exercises at the end of each section, and relevant chapter problems at the end of each chapter. A new Math Check allows quick access to the needed basic skill. The first chapter now includes brief introductions to several fundamental chemical concepts, and Chapter Synthesis Problems have been added to the end of each chapter to bring key concepts into one encompassing problem. Every concept in the text is clearly illustrated with one or more step-by-step examples. Making it Real essays have been updated to present timely and engaging real-world applications, emphasizing the relevance of the material they are learning. This edition continues the end-of-chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter. *Study Guide and Partial Solutions Manual, Fundamentals of General, Organic, and Biological Chemistry* Susan McMurry, 1996. This internationally acclaimed detective series is just the thing for lovers of those Number One Ladies looking for a darker, more realistic view of Botswana. Sue Baker Publishing News.

CliffsNotes AP Chemistry Bobrow Test Preparation Services, 2009-02-09. The book itself contains chapter-length subject reviews on every subject tested on the AP Chemistry exam, as well as both sample multiple-choice and free-response questions at each chapter's end. Two full-length practice tests with detailed answer explanations are included in the book.

Study Guide and Solutions Manual, Fundamentals of General, Organic, and Biological Chemistry, Third Edition John McMurry, Susan McMurry, 1999. Provides worked-out solutions to text problems along with chapter-by-chapter outlines and a variety of self-tests at the end of each chapter. *Chemistry, Student Study Guide* James E. Brady, Fred Senese, 2008-01-28. The image on the front cover depicts a carbon nanotube emerging from a glowing plasma of hydrogen and carbon as it forms around particles of a metal catalyst. Carbon nanotubes are a recently discovered allotrope of carbon. Three other allotropes of carbon—buckyballs, graphite, and diamond—are illustrated at the left, as is the molecule methane (CH_4) from which nanotubes and buckyballs can be made. The element carbon forms an amazing number of compounds with structures that follow from simple methane found in natural gas to the complex macromolecules that serve as the basis of life on our planet. The study of chemistry also follows from the simple to the more complex, and the strength of this text is that it enables students with varied backgrounds to proceed together to significant levels of achievement. *Student Solutions Manual for Introduction to General, Organic, and Biochemistry* Morris Hein, 2001-07-21. Experienced authors and educators Hein, Pattison, Best, and Arena recognize that both science and mathematics can be daunting subjects. They skillfully anticipate areas of difficulty and pace the text accordingly. Particular emphasis is placed on the understanding of how chemical principles relate to the readers.

lives and future careers **Chemistry Jeopardy** Brian Pressley, 2002 A quiz game designed to reinforce key chemistry concepts ideas and facts *Concepts and Applications in Chemistry for JEE Mains*, 2024-08-29 This book is strictly as per NCERT with deeper understanding of concepts This book covers wider application of concepts with focus on JEE Mains The book contains about 300 handpicked JEE Mains questions for MCQ practice on covering each topic separately The book has hundreds of topic wise illustrations and solved examples The book has hundreds of topic and subtopic wise practice questions The book has actual class like presentation of concepts The book is aimed at self paced learning **Study Guide with Selected Solutions** Karen Timberlake, 2002-08-06 Written by the author the Study Guide is keyed to the learning goals in the text and designed to promote active learning through a variety of exercises with answers and mastery exams Also contains complete solutions to odd numbered problems **Problem Exercises for General Chemistry** George Gilbert Long, Forrest C. Hentz, 1982 *Study Guide and Solutions Manual for Seager/Slabaugh's Chemistry for Today* Seager/Slabaugh, 2004 The fifth edition of the Study Guide and Student Solutions Manual has been updated to reflect all of the changes to the text This ancillary tests the student on the learning objectives in each chapter and provides answers to all of the even numbered end of chapter exercises New additional activities have been added to include a review of each section of the chapter and a section entitled Tying It All Together with a Laboratory Application **5 Steps to a 5 AP Chemistry, 2010-2011 Edition** John Moore, Richard H. Langley, 2010-01-01 A Perfect Plan for the Perfect Score We want you to succeed on your AP exam That s why we ve created this 5 step plan to help you study more effectively use your preparation time wisely and get your best score This easy to follow guide offers you a complete review of your AP course strategies to give you the edge on test day and plenty of practice with AP style test questions You ll sharpen your subject knowledge strengthen your thinking skills and build your test taking confidence with Full length practice exams modeled on the real test All the terms and concepts you need to know to get your best score Your choice of three customized study schedules so you can pick the one that meets your needs The 5 Step Plan helps you get the most out of your study time Step 1 Set Up Your Study Program Step 2 Determine Your Readiness Step 3 Develop the Strategies Step 4 Review the Knowledge Step 5 Build Your Confidence Topics include Reactions and Periodicity Stoichiometry Gases Thermodynamics Spectroscopy Light and Electrons Bonding Solids Liquids and Intermolecular Forces Solutions and Colligative Properties Kinetics Equilibrium Electrochemistry Nuclear Chemistry and Organic Chemistry Also includes AP Chemistry practice exams AP Advanced Placement Program and College Board are registered trademarks of the College Entrance Examination Board which was not involved in the production of and does not endorse this product *5 Steps to a 5 AP Chemistry, 2012-2013 Edition* Richard H. Langley, John Moore, 2011-06-10 A Perfect Plan for the Perfect Score We want you to succeed on your AP exam That s why we ve created this 5 step plan to help you study more effectively use your preparation time wisely and get your best score This easy to follow guide offers you a complete review of your AP course strategies to give you the edge on test day and plenty of practice

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A Guide to Modern Chemistry John F. Baxter, Jay A. Young, 1959 *Problem Exercises for General Chemistry* G. Gilbert Long, Forrest C. Hentz, 1986-08-14 A question answer supplement to accompany a 1 year general chemistry course This text is a multiple choice format and can be used with any standard general chemistry text The exercises emphasize the importance of problem solving and equation writing in the style used in general chemistry examinations and professional school aptitude examinations such as the MCAT and VCAT

Encyclopaedia Britannica, 1929 **The Encyclopaedia Britannica** James Louis Garvin, Franklin Henry Hooper, Warren E. Cox, 1929 *General, Organic, and Biochemistry Study Guide* Marcia L. Gillette, Ira Blei, Wendy Gloffke, George Odian, 2006-01-27 This study guide provides reader friendly reinforcement of the concepts covered in the textbook Features include Chapter outlines Are you able to Worked text problems Fill ins Test yourself Concept maps Can also be used for Blei and Odian s Organic and Biochemistry

Robinson Chemistry Study Guide Robinson, 1992 **Student Study Guide and Solutions Manual to accompany General Organic and Biological Chemistry, 1e** Kenneth W. Raymond, 2005-10-07 Finally readers have a shorter less intimidating introduction to general organic and biological chemistry Not only is Raymond s text concise it also takes an integrated approach to presenting important topics in a way that makes the material easier to understand In this approach similarities can be exploited and concepts reinforced The result is that readers see the strong connections that exist between these three branches of chemistry

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correct answer a explanation to test positive in the biuret test and negative in the sudan red test a substance would need to have protein and not have lipids fats milk is a good source of protein and skim milk is fat free so that is the answer wheat bread does not have protein or lipids meatloaf has both protein and lipids

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