

## ChemThink: Molecular Shapes

Objective:




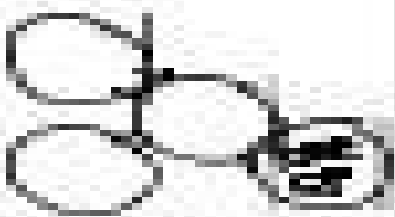

Lesson 8

Electrons

Page

Electron dot

Understanding that as the number of molecules that have lone pairs on an atom, the electron dot structure will change and each atom will have repulsion and it affects the shape of the molecule.

Number of lone pairs on atom	Electron dot structure	Bond Angles	Shape	E.g. Lewis dot structure
0		$180^\circ$	Linear	$H-H$
1		$120^\circ$	Trigonal Planar	
2		$180^\circ$	Bent	

# Molecular Shapes Chemthink Answns

**RD Boyd**



## **Molecular Shapes Chemthink Answr:**

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