

mind. This idea of gas molecules hitting the wall will be used often. When the generalized variable of pressure is discussed, the symbol P is used.

IV. Amount of Gas

The amount of gas present is measured in moles (symbol = mol) or in grams (symbol = g or gm). Typically, if grams are used, you will need to convert to moles at some point. When the generalized variable of amount in moles is discussed, the letter "n" is used as the symbol (note: the letter is in lowercase. Don't use uppercase.).

3. What is Gas Laws?

The **gas laws** have been around to assist scientists in finding volumes, amount, pressures and temperature when coming to matters of gas. The gas laws consist of three primary laws: Charles' Law, Boyle's Law and Avogadro's Law (all of which will later combine into the General Gas Equation and Ideal Gas Law).

Sample Exercises

1. A gas occupies 11.2 liters at 0.860 atm. What is the pressure if the volume becomes 15.0 L?

The image shows a handwritten solution on lined paper. At the top left, there is a small box containing the number 1. The solution is organized into three steps. Step 1 lists the initial conditions: $P_1 = 0.860 \text{ atm}$, $P_2 = ?$, $V_1 = 11.2 \text{ L}$, and $V_2 = 15.0 \text{ L}$. Step 2 shows the equation $\frac{P_1 V_1}{V_2} = \frac{P_2 V_2}{V_2} \rightarrow \frac{P_1 V_1}{V_2} = P_2$. Step 3 shows the calculation for P_2 : $P_2 = \frac{P_1 V_1}{V_2}$, followed by $= \frac{0.860 \text{ atm} (11.2 \text{ L})}{15.0 \text{ L}}$, and finally $[= 0.64 \text{ atm}]$. A small box with the number 2 is at the bottom right of the calculation.

$$\begin{aligned} \text{Step 1: } & P_1 = 0.860 \text{ atm} \\ & P_2 = ? \\ & V_1 = 11.2 \text{ L} \\ & V_2 = 15.0 \text{ L} \\ \text{Step 2: } & \frac{P_1 V_1}{V_2} = \frac{P_2 V_2}{V_2} \rightarrow \frac{P_1 V_1}{V_2} = P_2 \\ \text{Step 3: } & P_2 = \frac{P_1 V_1}{V_2} \\ & = \frac{0.860 \text{ atm} (11.2 \text{ L})}{15.0 \text{ L}} \\ & [= 0.64 \text{ atm}] \end{aligned}$$

Molecular Composition Of Gases Answers

Judd E. Hollander



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