

## Gas Law Formula

Gas Law	Formula	Description
Boyle's Law	$P_1V_1 = P_2V_2$	At constant $T$ , as pressure increases, volume decreases.
Charles' Law	$\frac{V_1}{T_1} = \frac{V_2}{T_2}$	At constant $P$ , as volume increases, temperature increases.
Gay-Lussac's Law	$\frac{P_1}{T_1} = \frac{P_2}{T_2}$	At constant $V$ , as pressure increases, temperature increases.
Combined Law	$\frac{P_1V_1}{T_1} = \frac{P_2V_2}{T_2}$	Obtained by combining Boyle's Law, Charles' Law and Gay-Lussac's Law.
Ideal Gas Law	$PV = nRT$	
$V$ = volume in $\text{dm}^3$ $T$ = temperature in K		$P$ = pressure in kPa $n$ = number of moles $R$ = ideal gas constant

# Mixed Gas Law Calculation Work For Problems

**Tao Wei**



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