

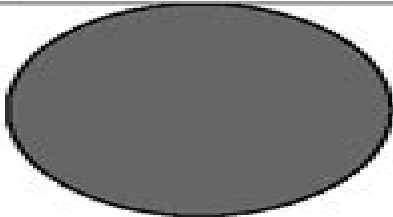


# Atomic Theory Worksheet

*Objective: You will create an abbreviated atomic theory timeline to show how the scientists involved in the atomic theory worked together to construct the modern model of the atom.*

## Part 1 – Atomic Theory Timeline

Choose 3 scientists from the options below and complete the table.

Options: John Dalton, Niels Bohr, JJ Thomson, Erwin Schrödinger, Ernest Rutherford

Content	Scientist 1	Scientist 2	Scientist 3
Timeframe [in order]	1803	1897	1911
Scientist Name	John Dalton	J.J. Thompson	Ernest Rutherford
What did he contribute to the atomic theory?	He contributed that atoms can not be made or destroyed, a chemical reaction only can happen when atoms are rearranged, as well all atoms of a specific element are the same in mass and properties.	He contributed that atoms have charged particles that are called electrons, he also thought that matter is electronically neutral, he also thought that electrons are dispersed evenly through an atom.	He contributed that an atom has a very small positively charged center called the nucleus and it contains most of the atoms mass. He also added that atoms have a cloud of electrons that circle around the nucleus.
Draw a sketch of their model [by hand or using the drawing tool] and label all parts.		 atom electrons	 electrons nucleus

## Part 2 – Reflection Questions

1. Describe two details of the modern atomic model.

Atoms have a nucleus of protons and neutrons and a cloud around the nucleus that contains the electrons.

2. Explain how the atomic model has changed over time.

The Rutherford model got replaced by the Bohr model because scientists found out that electrons have a specific orbit around the nucleus.

# Modern Atomic Theory Concept Review Answers

**Morris Hein, Scott Pattison, Susan  
Arena, Leo R. Best**



## Modern Atomic Theory Concept Review Answers:

Basic Concepts of Chemistry, 9e Study Guide and Solutions Manual Leo J. Malone, Theodore O. Dolter, 2012-01-03 The 9th edition of Malone's Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment. New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections, assessment exercises at the end of each section, and relevant chapter problems at the end of each chapter. A new Math Check allows quick access to the needed basic skill. The first chapter now includes brief introductions to several fundamental chemical concepts, and Chapter Synthesis Problems have been added to the end of each chapter to bring key concepts into one encompassing problem. Every concept in the text is clearly illustrated with one or more step-by-step examples. Making it Real essays have been updated to present timely and engaging real-world applications, emphasizing the relevance of the material they are learning. This edition continues the end-of-chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter.

**Merrill Chemistry** Robert C. Smoot, Richard G. Smith, Jack Price, 1998

**Foundations of College Chemistry** Morris Hein, Susan Arena, Cary Willard, 2023 Foundations of College Chemistry 16th edition presents chemistry as a modern vital subject and is designed to make introductory chemistry accessible to all beginning students. It is intended for students who have never taken a chemistry course or those who had a significant interruption in their studies but plan to continue with the general chemistry sequence. The central focus is to make chemistry interesting and understandable and teach students the problem-solving skills they will need. This International Adaptation offers new and updated content with improved presentation of all course material. It builds on the strengths of previous editions, including clear explanations and step-by-step problem solving. The material emphasizes real-world applications of chemistry as the authors develop the principles that form the foundation for the further study of chemistry. There is new and expanded coverage of polarizing power and polarizability, Fajans' rules, collision number, and mean free path, abnormal molecular masses, and van't Hoff factor and applications of radioactivity.

*Study Guide to accompany Basic Concepts of Chemistry, 7th Edition* Leo J. Malone, 2003-02-20 Work more effectively and gauge your progress along the way. This Study Guide that is designed to accompany Malone's Chemistry 7th Edition includes chapter summaries, new terms, self-tests, answers to self-tests, and solutions to selected problems. This easy-to-read introduction presents chemistry as a living relevant science. Chemistry 7th Edition encourages critical thinking and helps readers overcome the math difficulties that often prevent them from developing a full understanding of the subject. This new seventh edition builds on its core strengths of pedagogy, driving the connections between ideas, mathematics in context, not just an appendix, and an extensive problem-solving emphasis with an updated design and more molecular art. In addition, the seventh edition expands its applications and online options. One of the briefest books in the market, it still provides sufficient depth for the basic concepts of chemistry.

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*Foundations of Chemistry* Morris Hein,Susan Arena,1996 A textbook introducing matter atomic theory ionization and other aspects of chemistry to the high school student      **Concepts of Electricity** Lawrence P. Larsen,1988      *The Quarterly Review of Biology* Raymond Pearl,1929 Includes section New biological books and other bibliographies      **Chemical Interactions** ,2005      The Review of Reviews William Thomas Stead,1898      **Holt Physical Science** Mapi M. Cuevas,1994

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