

Chemistry Semester 1 Final Exam Review

- 1.
 2. What is a definition of Chemistry?
 3. Matter includes all of the following EXCEPT:
 - a. air
 - b. light
 - c. smoke
 - d. water vapor
 4. What is matter?
 5. A physical change occurs when a
 - a. peach spoils
 - b. bracelet turns your wrist green
 - c. copper bowl tarnishes
 - d. glue gun melts a glue stick
 6. The particles in a solid are
 - a. packed closely together
 - b. very far apart
 - c. constantly in motion
 - d. able to slide past one another
- Of the following groups, 1, 2, 14, 15, 17, 18:
7. Group(s) _____ contains only metals
 8. Group (s) _____ is called a Noble gases
 9. Group(s) _____ has 7 electrons on the valence shell
 10. Group(s) _____ needs to gain 3 electrons to fill a stable octet
 11. Group(s) _____ has a 2+ oxidation state
 12. Group(s) _____ has an electron dot notation of $\bullet X \bullet$
 13. Using the periodic table, an element similar to carbon would be _____
 14. Based on their location on the periodic table, you could infer that _____ are very unreactive
 15. The horizontal row on the periodic table is called a(n) _____
 16. The vertical column of blocks on the periodic table is called a(n) _____
 17. The elements on the zigzag line in the periodic table are _____
 18. Which is NOT a property of metal?
 - a. malleability
 - b. ability to lose electrons
 - c. unreactivity
 - d. ability to conduct heat
 19. A volume of 1 cc (cm^3) is the same as _____ mL
 20. To determine density, the quantities that must be measured are _____ and _____
 21. The relationship between the mass m of a material, its volume V , and its density D is
 - a. $V = mD$
 - b. $V/m = D$
 - c. $DV = m$
 - d. $D + V = m$
 22. The density of aluminum is 2.70 g/cm^3 . The volume of a solid piece of Aluminum is 1.50 cm^3 . Find the mass
 23. The density of pure diamond is 3.5 g/cm^3 . The mass of a diamond is 0.25 g . Find the volume
 24. What statement about density is true?
 - a. two samples of a pure substance may have different densities if they are different sizes
 - b. the density of a sample depends on its location on Earth
 - c. a cylinder is always used to measure the volume
 - d. density is a physical property that remains constant for a pure substance
 25. The number of grams equal to 0.5 kg is _____
 26. In a graph the _____ variable is placed on the x axis and the _____ variable is placed on the y axis.
 27. The variable that you change in the lab is the _____ variable. The variable that responds to the change is called the _____ variable. All other things do not change and are called _____
 28. At first Dalton thought the atom was _____, then Rutherford discovered the _____ of the atom, then the subatomic particles _____, _____, and _____ were discovered
 29. The positively charged particle is the _____ found in the _____ of the atom
 30. The subatomic particle that has about the same mass as the proton, but with no electrical charge, is called a _____ and is found in the _____ of the atom
 31. The _____ is negatively charged and is found in the _____ of the atom
 32. _____ is the person credited with placing electrons on levels and _____ established the electron cloud theory.
 33. The forces that hold the particles together in the nucleus are _____ forces
 34. Isotopes are atoms of the same element that _____
 35. All isotopes of hydrogen contain _____
 36. What is the atomic number for Nitrogen? To what subatomic particle does this correspond?
 37. Zinc -66 has _____ protons and _____ neutrons
 38. What is the electron configuration for sodium? Chlorine? Neon?
 39. What is the dot notation for magnesium? Bromine? Argon?
 40. How many electrons is each element trying to achieve in the valence level?
 41. Which family is entirely radioactive?
 42. Which family will gain two electrons to complete their octet?
 43. Where are the transition metals located? What is special about their electron configuration that gives them special properties?
 44. As you move left to right across a period on the table the size (radius) _____

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