

Modern Analytical Chemistry Worksheet

1. What is the relationship between absorbance and transmittance? (3 points)

Transmittance is the fraction of radiation (from the source) that passes through the sample. It is given by the formula $T = P/P_0$. Meanwhile, absorbance is the intensity of radiation absorbed by the sample and it is the negative logarithm of transmittance ($A = -\log T$).

2. A solution containing the complex formed between Bi (III) and thiourea has a molar absorptivity of $9.32 \times 10^3 \text{ L/cm}\cdot\text{mol}$ at 470nm.

(a) What is the absorbance of a $5.67 \times 10^{-5} \text{ M}$ solution of the complex at 470 nm in a 1.00 cm-cell?

$$A = \epsilon bc$$

Given:

$$\epsilon = 9.32 \times 10^3 \text{ L/cm}\cdot\text{mol}$$

$$b = 1.00 \text{ cm}$$

$$c = 5.67 \times 10^{-5} \text{ M}$$

$$A = ?$$

$$\text{Solution: } A = (9.32 \times 10^3) \times (1.00 \text{ cm}) \times (5.67 \times 10^{-5}) = 0.528$$

(b) What is the percent transmittance of the solution described in (a)

$$A = -\log T$$

Given:

$$A = 0.528$$

$$\%T = ?$$

Solution:

$$-\log T = 0.528$$

$$T = (10^{0.528})^{-1} = 0.296483139$$

$$\%T = 0.296483139 \times 100\% = 29.6\%$$

(c) What is the molar concentration of the complex in a solution that has the absorbance described in (a) when measured at 470 nm in a 2.50 cm cell?

$$A = \epsilon bc$$

Given:

$$A = 0.528$$

$$\epsilon = 9.32 \times 10^3 \text{ L/cm}\cdot\text{mol}$$

$$b = 2.50 \text{ cm}$$

$$c = ?$$

Solution:

$$c = \frac{0.528}{(9.32 \times 10^3 \text{ L/cm}\cdot\text{mol})(2.50 \text{ cm})} = 2.27 \times 10^{-5} \text{ M}$$

Modern Chemistry Ch 13 Review Answers

Daniel F McAuley



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Modern Chemistry Holt Rinehart & Winston, Holt, Rinehart and Winston Staff, 2001 Modern Enolate Chemistry Manfred Braun, 2016-02-23 Authored by one of the world's leading synthetic chemists in the field this reference presents modern enolate chemistry with an emphasis on metal O enolates in asymmetric synthesis While great care is taken to cover novel successful concepts such classical methods as the famous Evans enolates are equally highlighted Throughout the book representative reaction procedures are presented thus helping readers to find the best solution for their own synthetic problem Of high interest to synthetic chemists in academia as well as the pharmaceuticals agrochemicals and fine chemicals industries **Foundations of College Chemistry, Student Solutions Manual** Morris Hein, Susan Arena, Cary

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