

Modern Analytical Chemistry Worksheet

1. What is the relationship between absorbance and transmittance? (3 points)

Transmittance is the fraction of radiation (from the source) that passes through the sample. It is given by the formula $T = P/P_0$. Meanwhile, absorbance is the intensity of radiation absorbed by the sample and it is the negative logarithm of transmittance ($A = -\log T$).

2. A solution containing the complex formed between Bi (III) and thiourea has a molar absorptivity of $9.32 \times 10^3 \text{ L/cm}\cdot\text{mol}$ at 470nm.

(a) What is the absorbance of a $5.67 \times 10^{-5} \text{ M}$ solution of the complex at 470 nm in a 1.00 cm-cell?

$$A = \epsilon bc$$

Given:

$$\epsilon = 9.32 \times 10^3 \text{ L/cm}\cdot\text{mol}$$

$$b = 1.00 \text{ cm}$$

$$c = 5.67 \times 10^{-5} \text{ M}$$

$$A = ?$$

$$\text{Solution: } A = (9.32 \times 10^3) \times (1.00 \text{ cm}) \times (5.67 \times 10^{-5}) = 0.528$$

(b) What is the percent transmittance of the solution described in (a)

$$A = -\log T$$

Given:

$$A = 0.528$$

$$\%T = ?$$

Solution:

$$-\log T = 0.528$$

$$T = (10^{0.528})^{-1} = 0.296483139$$

$$\%T = 0.296483139 \times 100\% = 29.6\%$$

(c) What is the molar concentration of the complex in a solution that has the absorbance described in (a) when measured at 470 nm in a 2.50 cm cell?

$$A = \epsilon bc$$

Given:

$$A = 0.528$$

$$\epsilon = 9.32 \times 10^3 \text{ L/cm}\cdot\text{mol}$$

$$b = 2.50 \text{ cm}$$

$$c = ?$$

Solution:

$$c = \frac{0.528}{(9.32 \times 10^3 \text{ L/cm}\cdot\text{mol})(2.50 \text{ cm})} = 2.27 \times 10^{-5} \text{ M}$$

Modern Chemistry Chapter 13 Mixed Review Answers

Padhraic Smyth



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Table of Contents Modern Chemistry Chapter 13 Mixed Review Answers

1. Understanding the eBook Modern Chemistry Chapter 13 Mixed Review Answers
 - The Rise of Digital Reading Modern Chemistry Chapter 13 Mixed Review Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Modern Chemistry Chapter 13 Mixed Review Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Modern Chemistry Chapter 13 Mixed Review Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Modern Chemistry Chapter 13 Mixed Review Answers
 - Personalized Recommendations
 - Modern Chemistry Chapter 13 Mixed Review Answers User Reviews and Ratings
 - Modern Chemistry Chapter 13 Mixed Review Answers and Bestseller Lists

5. Accessing Modern Chemistry Chapter 13 Mixed Review Answers Free and Paid eBooks
 - Modern Chemistry Chapter 13 Mixed Review Answers Public Domain eBooks
 - Modern Chemistry Chapter 13 Mixed Review Answers eBook Subscription Services
 - Modern Chemistry Chapter 13 Mixed Review Answers Budget-Friendly Options
6. Navigating Modern Chemistry Chapter 13 Mixed Review Answers eBook Formats
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 - Modern Chemistry Chapter 13 Mixed Review Answers Compatibility with Devices
 - Modern Chemistry Chapter 13 Mixed Review Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Modern Chemistry Chapter 13 Mixed Review Answers
 - Highlighting and Note-Taking Modern Chemistry Chapter 13 Mixed Review Answers
 - Interactive Elements Modern Chemistry Chapter 13 Mixed Review Answers
8. Staying Engaged with Modern Chemistry Chapter 13 Mixed Review Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Modern Chemistry Chapter 13 Mixed Review Answers
9. Balancing eBooks and Physical Books Modern Chemistry Chapter 13 Mixed Review Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Modern Chemistry Chapter 13 Mixed Review Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Modern Chemistry Chapter 13 Mixed Review Answers
 - Setting Reading Goals Modern Chemistry Chapter 13 Mixed Review Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Modern Chemistry Chapter 13 Mixed Review Answers
 - Fact-Checking eBook Content of Modern Chemistry Chapter 13 Mixed Review Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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