

# CHEMICAL EQUILIBRIUM



## EQUILIBRIUM CONSTANT 'K'.

For a General Reaction  $aA + bB \rightleftharpoons pP + qQ$

The **equilibrium constant** expression is

$$K_c = \frac{[P]^p [Q]^q}{[A]^a [B]^b}$$

where  $K_c$  is the **Equilibrium Constant** (or  $K_p$  if they are all gases)

## RELATION BETWEEN $K_p$ AND $K_c$

For the Reaction  $aA + bB \rightleftharpoons cC + dD$

$$K_p = \frac{[P_c]^c \times [P_D]^d}{[P_A]^a \times [P_B]^b} = \frac{[C]^c [D]^d}{[A]^a [B]^b} \frac{(RT)^{c+d}}{(RT)^{a+b}}$$

$$K_p = K_c (RT)^{\Delta n_g}$$

$$\text{if } \Delta n_g = 0 \implies K_p = K_c$$

Where,  $\Delta n_g = (c+d) - (a+b)$

= no. of moles of gaseous products - no. of moles of gaseous Reactants

## WHAT DOES THE VALUE OF 'K' MEAN ?

(a)  $K \gg 1$

Reactants



Products

If  $K \gg 1$ , the reaction is product-favoured; product predominates at **Equilibrium**.

(a)  $K \ll 1$

Reactants



Products

If  $K \ll 1$ , the reaction is reactant-favoured; reactant predominates at **Equilibrium**.

(a)  $K = 1$

The reaction lies in the middle (mix of reactants and products)

## MAGNITUDE OF 'K'.

Small ( $K < 10^{-3}$ )



Reactants



Products

Mostly Reactants

Intermediate ( $10^{-3} \leq K \leq 10^3$ )



Reactants



Products

Significant amounts of reactants and products

Large ( $K > 10^3$ )



Reactants



Products

Mostly Products

# Modern Chemistry Section Chemical Equilibrium

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