CHAPTER 9 REVIEW

Stoichiometry

MIXED REVIEW

SHORT ANSWER Answer the following questions in the space provided.

Given the following equation: C₂H_d(g) + xO₂(g) → 3CO₂(g) + 2H₂O(g)

a. What is the value of the coefficient x in this equation?

b. What is the molar mass of C₃H₄?

c. What is the mole ratio of O₂ to H₂O in the above equation?

d. How many moles are in an 8.0 g sample of C₃H₄?

e. If z mol of C₃H₄ react, how many moles of CO₂ are produced, in terms of z?

2. a. What is meant by ideal conditions relative to stoichiometric calculations?

b. What function do ideal stoichiometric calculations serve?

c. Are actual yields typically larger or smaller than theoretical yields?

PROBLEMS Write the answer on the line to the left. Show all your work in the space provided.

3. Assume the reaction represented by the following equation goes all the way to completion:

$$N_2 + 3H_2 \rightarrow 2NH_3$$

a. If 6 nsol of H₂ are consumed, how many moles of NH₃ are produced?

b. How many grams are in a sample of NH₃ that contains 3.0 × 10²³ molecules?

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David Baud

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