

Stoichiometry Worksheet # 1

- 1) Glucose is used as a source of energy by the human body. The overall reaction in the body is



Calculate the number of moles of oxygen needed to oxidize 12.5 g of $\text{C}_6\text{H}_{12}\text{O}_6$.

- 2) Ammonia is synthesized from hydrogen and nitrogen according to the following equation.



If an excess of nitrogen is reacted with 3.41 moles of hydrogen gas, how many grams of ammonia (NH_3) can be produced?

- 3) Assume that in the decomposition of potassium chlorate, KClO_3 ,



80.5 moles of O_2 form. How many grams of potassium chlorite, the other product would be formed?

- 4) In a single displacement reaction,



9.23 moles of aluminum react with excess hydrochloric acid. How many moles of hydrogen will be produced?

- 5) The compound "Cisplatin" $\text{PtCl}_2(\text{NH}_3)_2$, has been found to be effective in treating some types of cancer. It can be synthesized using the following reaction



A. How many moles of "cisplatin" can be produced from 2.50 g K_2PtCl_4 ?

B. How many moles of NH_3 would be needed if 3.8 moles of K_2PtCl_4 were present?

Modern Chemistry Answer Key Chapter 9 Stoichiometry

Wade A. Freeman



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Modern Chemistry Raymond E. Davis, 1999 2000 2005 State Textbook Adoption Rowan Salisbury *Modern Chemistry* Nicholas D. Tzimopoulos, 1993 *Principles of Modern Chemistry* David W. Oxtoby, H. Pat Gillis, Laurie J. Butler, 2016-01-01 Long considered the standard for honors and high level mainstream general chemistry courses PRINCIPLES OF MODERN CHEMISTRY continues to set the standard as the most modern rigorous and chemically and mathematically accurate text on the market This authoritative text features an atoms first approach and thoroughly revised chapters on Quantum Mechanics and Molecular Structure Chapter 6 Electrochemistry Chapter 17 and Molecular Spectroscopy and Photochemistry Chapter 20 In addition the text utilizes mathematically accurate and artistic atomic and molecular orbital art and is student friendly without compromising its rigor End of chapter study aids focus on only the most important key objectives equations and concepts making it easier for students to locate chapter content while applications to a wide range of disciplines such as biology chemical engineering biochemistry and medicine deepen students understanding of the relevance of chemistry beyond the classroom Chemistry: Core Concepts, 3rd Edition Allan Blackman, Daniel Southam, Gwendolyn Lawrie, Natalie Williamson, Christopher Thompson, 2024-01-09 The third edition of Chemistry Core Concepts Blackman et al has been developed by a group of leading chemistry educators for students entering university with little or no background in chemistry Available as a full colour printed textbook with an interactive eBook code this title enables every student to master concepts and succeed in assessment Lecturers are supported with an extensive and easy to use teaching and learning package **Introductory Chemistry** Steven S. Zumdahl, 1996 **Ligand-Binding Basics** Jannette Carey, 2025-11-10 A concise and accessible textbook covering ligand binding theory in chemistry biology and drug development In Ligand binding Basics Evaluating Intermolecular Affinity Specificity Stoichiometry and Cooperativity accomplished chemist Professor Jannette Carey introduces ligand binding in a thorough and practical way for those new to the topic as well as anyone seeking a connection between theory and experiment Using a minimum of mathematical formalism this book offers analytical rigor while remaining accessible to non specialist practitioners It provides readers with the skills they need to analyze their own binding data or published results helping them develop an intuitive grasp of ligand binding phenomena integrated with structural and thermodynamic understanding Topics covered include Application of the principles of equilibrium mass action and mass balance to derive the basic equations that describe all binding processes Recommended approaches for plotting and graphical analysis of binding data Strategies for designing analyzing interpreting and troubleshooting experiments from the perspective of ligand binding theory Review of selected examples that illustrate integration of structural and thermodynamic analysis Perfect for students and educators in chemistry biochemistry molecular biology and pharmaceutical science Ligand binding Basics will also appeal to practitioners who aim to study ligand binding in any molecular system General Chemistry with Qualitative Analysis Ralph H. Petrucci, Robert K. Wismer, 1987 Modern University Chemistry,

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