

Chapter 10 test

Do I want to be a manager?

Match each item with the correct statement below.

- | | |
|----------------------------|--------------------------------------|
| a. representative particle | d. percent composition |
| b. mole | e. standard temperature and pressure |
| c. Avogadro's number | f. empirical formula |

1. the number of representative particles of a substance present in 1 mole of that substance
2. an atom, an ion, or a molecule, depending upon the way a substance commonly exists
3. the SI unit used to measure amount of substance
4. 0°C and 1 atm
5. the percent by mass of each element in a compound
6. the smallest whole number ratio of the atoms in a compound

Multiple Choice

Identify the letter of the choice that best completes the statement or answers the question.

7. What SI unit is used to measure the number of representative particles in a substance?
- a. kilogram
b. ampere
c. kelvin
d. mole
8. How many hydrogen atoms are in 5 molecules of isopropyl alcohol, C₃H₇O?
- a. $5 \times (6.02 \times 10^{23})$
b. 5
c. 35
d. $35 \times (6.02 \times 10^{23})$
9. All of the following are equal to Avogadro's number EXCEPT _____.
- a. the number of atoms of bromine in 1 mol Br₂
b. the number of atoms of gold in 1 mol Au
c. the number of molecules of nitrogen in 1 mol N₂
d. the number of molecules of carbon monoxide in 1 mol CO
10. How many moles of tungsten atoms are in 4.8×10^{25} atoms of tungsten?
- a. 8.0×10^{-2} moles
b. 8.0×10^1 moles
c. 1.3×10^{-4} moles
d. 1.3×10^{-2} moles
11. How many atoms are in 0.075 mol of titanium?
- a. 1.2×10^{23}
b. 2.2×10^{24}
c. 6.4×10^{23}
d. 4.5×10^{22}
12. How many molecules are in 2.10 mol CO₂?
- a. 2.53×10^{24} molecules
b. 3.79×10^{24} molecules
c. 3.49×10^{-24} molecules
d. 1.26×10^{24} molecules

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John C. Flynn



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Table of Contents Modern Chemistry Chapter 10 Test

1. Understanding the eBook Modern Chemistry Chapter 10 Test
 - The Rise of Digital Reading Modern Chemistry Chapter 10 Test
 - Advantages of eBooks Over Traditional Books
2. Identifying Modern Chemistry Chapter 10 Test
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Modern Chemistry Chapter 10 Test
 - User-Friendly Interface
4. Exploring eBook Recommendations from Modern Chemistry Chapter 10 Test
 - Personalized Recommendations
 - Modern Chemistry Chapter 10 Test User Reviews and Ratings
 - Modern Chemistry Chapter 10 Test and Bestseller Lists
5. Accessing Modern Chemistry Chapter 10 Test Free and Paid eBooks
 - Modern Chemistry Chapter 10 Test Public Domain eBooks
 - Modern Chemistry Chapter 10 Test eBook Subscription Services
 - Modern Chemistry Chapter 10 Test Budget-Friendly Options

6. Navigating Modern Chemistry Chapter 10 Test eBook Formats
 - ePub, PDF, MOBI, and More
 - Modern Chemistry Chapter 10 Test Compatibility with Devices
 - Modern Chemistry Chapter 10 Test Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Modern Chemistry Chapter 10 Test
 - Highlighting and Note-Taking Modern Chemistry Chapter 10 Test
 - Interactive Elements Modern Chemistry Chapter 10 Test
8. Staying Engaged with Modern Chemistry Chapter 10 Test
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Modern Chemistry Chapter 10 Test
9. Balancing eBooks and Physical Books Modern Chemistry Chapter 10 Test
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Modern Chemistry Chapter 10 Test
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Modern Chemistry Chapter 10 Test
 - Setting Reading Goals Modern Chemistry Chapter 10 Test
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Modern Chemistry Chapter 10 Test
 - Fact-Checking eBook Content of Modern Chemistry Chapter 10 Test
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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