3 NaOH + AI (NO3)3(2) -> AI (OH)3(5) + 3 Na NO3(24) What is the maximum mass of aluminum hydroxide that can be produced when 25.0 mL of 0.248 M sodium hydroxide solution are mixed with 25.0 mL of 0.189 M aluminum nitrate solution? Find mol of each reactant 0.0250 L x 0.248 mol = \$46.2 x10-3 mol NaOH need 3 times as much Na0 0.0250 L x 0.189 mol = 4.73 x103 mol Al (NO3) 3 as AI (NO3) 2 but we only have 6.20 × 10-3 mol NaOH x 1 mol AI(OH) x 78.0 9 + 1 (OH) 3 = 1.3 x as much NaOH is limiti. 0.161 9 A1 (OH)3 8. Indicate which of the following are redox reactions. For those that are redox reactions, label the reactant that is the oxidizing agent and the reactant that is the reducing agent. You do not have to balance these reactions! $C_2H_6 + O_2 \rightarrow CO_2 + H_2O$ 28. Redox agent agent $Li_2CO_3 + H_2O$ b. Not redox +5 -1 PC1₂ G. agent agent Aqueous dichromate ion reacts with aqueous cyanide ion (CN-) to produce chromium (III) ion and aqueous cyanate (CNO-) ion. Balance the equation using the half-reaction method, assuming the reaction occurs in acidic solution. Cr2032- + CN - - Cr3+ + CNO-6e" + 14H" + Cr2072" -> 2Cr3+ + 7H20 (H20+CN- - CNO- + 2H+ + 2e-)3 8 MH+ + Cr20+2- + 3H20 + 3CN- -> 2Cr3+ + 7H20+ 3CN0+ 6H4 10. Balance the following redox reaction in basic solution. → IO₃ (aq) ClO₁ (aq) (6e" + 6H++ C10," -> C1" + 3H,0)5 (GH20+ I2 -> 2 103 + 12H+ + 10e-)3 5 C103 + 3H20 + 3T2 - 5 C1 + 6 IO3 + 6 H+ 04 to each side: 6 04 + 5 C103 + 3 H20+ 3 I2 → 5 C1 + 6 IO3 + 6 H++60H) fral: 604- + 5010; + 3 T2 -> 501- + 6 IO; + 3 H2

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Thomas Engel

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