

# CHEMICAL BONDING

## Type of bonds

- Chemical bond formation takes place when  
(A) Energy is absorbed  
(B) Forces of attraction are overcome by the forces of repulsion  
(C) Forces of repulsion are overcome by forces of attraction  
(D) Forces of attraction are equal to the forces of repulsion
- Most favourable conditions for ionic bond formation are  
(A) Low charge ion ions, large cation, small anion  
(B) High charge on ions, large cation, small anion  
(C) High charge on ions, small cation, large anion  
(D) High charge on ions, small cation, large anion
- Which of the following shows the higher lattice energy?  
(A) RbF (B) CsF (C) NaF (D) KF
- Which of the following is least soluble in water?  
(A) BaF<sub>2</sub> (B) SiF<sub>4</sub> (C) CaF<sub>2</sub> (D) MgF<sub>2</sub>
- Which of the following sets does not contain any ionic species :  
(A) NH<sub>4</sub>Cl, OF<sub>2</sub>, H<sub>2</sub>S (B) CO<sub>2</sub>, Cl<sub>2</sub>, CCl<sub>4</sub> (C) BF<sub>3</sub>, AlF<sub>3</sub>, TiF<sub>3</sub> (D) I<sub>2</sub>, CaO, CH<sub>3</sub>Cl
- There is no S-S bond in :  
(A) S<sub>2</sub>O<sub>3</sub><sup>2-</sup> (B) S<sub>2</sub>O<sub>4</sub><sup>2-</sup> (C) S<sub>4</sub>O<sub>6</sub><sup>2-</sup> (D) S<sub>2</sub>O<sub>7</sub><sup>2-</sup>
- The species that does not contain peroxide ions are :  
(A) PbO<sub>2</sub> (B) H<sub>2</sub>O<sub>2</sub> (C) SrO<sub>2</sub> (D) BaO<sub>2</sub>
- Which of the following compounds contains electrovalent and covalent bonds ?  
(A) CS<sub>2</sub> (B) SO<sub>2</sub> (C) KCN (D) Phosphine
- In which solvent has maximum solubility  
(A) H<sub>2</sub>O (B) C<sub>2</sub>H<sub>5</sub>OH (C) CH<sub>3</sub>COCH<sub>3</sub> (D) C<sub>2</sub>H<sub>5</sub>OC<sub>2</sub>H<sub>5</sub>
- Which of the following is a non-directional bond  
(A) Metallic bond (B) Covalent bond (C) Co-ordinate bond (D) All are directional
- Find the number of σ and π bonds in the following compounds  
(A) CH<sub>3</sub>CH<sub>3</sub> (B) C<sub>6</sub>(CN)<sub>6</sub> (C) CH<sub>2</sub>=CH<sub>2</sub> (D) C<sub>2</sub>H<sub>2</sub>Cl
- Point out the inter particle force operating in each of the following :  
(A) H<sub>2</sub>O(l) (B) H<sub>2</sub>O(s) (C) Dryice (D) KCl  
(E) Hg (F) CF<sub>4</sub> (G) Rock salt (H) Diamond
- True / False  
(i) Graphite, a covalent molecule, conducts current in solid state  
(ii) Intra molecular H-Bonding decreases the solubility of compound in water

## Matching

	Column I		Column - II
(a)	H <sub>2</sub> O	(p)	Ionic bond
(b)	C <sub>2</sub> H <sub>2</sub>	(q)	Covalent bond
(c)	KCN	(r)	Co-ordinate bond
(d)	K <sub>4</sub> [Fe(CN) <sub>6</sub> ]	(s)	Hydrogen bond

## Fagan's rule, Physical property

- Amongst LiCl, BeCl<sub>2</sub>, MgCl<sub>2</sub> and RbCl the compounds with greatest and least ionic character respectively are :  
(A) LiCl and RbCl (B) RbCl and BeCl<sub>2</sub> (C) RbCl and MgCl<sub>2</sub> (D) MgCl<sub>2</sub> and BeCl<sub>2</sub>
- The formal charge of the O atoms in the ion  $[\text{O}=\text{N}=\text{O}]_{-2}$  is

# Modern Chemistry Chemical Bonding Section Quiz

## Answers

**B Lingard**



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## **Table of Contents Modern Chemistry Chemical Bonding Section Quiz Answers**

1. Understanding the eBook Modern Chemistry Chemical Bonding Section Quiz Answers
  - The Rise of Digital Reading Modern Chemistry Chemical Bonding Section Quiz Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Modern Chemistry Chemical Bonding Section Quiz Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Modern Chemistry Chemical Bonding Section Quiz Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Modern Chemistry Chemical Bonding Section Quiz Answers

- Personalized Recommendations
  - Modern Chemistry Chemical Bonding Section Quiz Answers User Reviews and Ratings
  - Modern Chemistry Chemical Bonding Section Quiz Answers and Bestseller Lists
5. Accessing Modern Chemistry Chemical Bonding Section Quiz Answers Free and Paid eBooks
    - Modern Chemistry Chemical Bonding Section Quiz Answers Public Domain eBooks
    - Modern Chemistry Chemical Bonding Section Quiz Answers eBook Subscription Services
    - Modern Chemistry Chemical Bonding Section Quiz Answers Budget-Friendly Options
  6. Navigating Modern Chemistry Chemical Bonding Section Quiz Answers eBook Formats
    - ePub, PDF, MOBI, and More
    - Modern Chemistry Chemical Bonding Section Quiz Answers Compatibility with Devices
    - Modern Chemistry Chemical Bonding Section Quiz Answers Enhanced eBook Features
  7. Enhancing Your Reading Experience
    - Adjustable Fonts and Text Sizes of Modern Chemistry Chemical Bonding Section Quiz Answers
    - Highlighting and Note-Taking Modern Chemistry Chemical Bonding Section Quiz Answers
    - Interactive Elements Modern Chemistry Chemical Bonding Section Quiz Answers
  8. Staying Engaged with Modern Chemistry Chemical Bonding Section Quiz Answers
    - Joining Online Reading Communities
    - Participating in Virtual Book Clubs
    - Following Authors and Publishers Modern Chemistry Chemical Bonding Section Quiz Answers
  9. Balancing eBooks and Physical Books Modern Chemistry Chemical Bonding Section Quiz Answers
    - Benefits of a Digital Library
    - Creating a Diverse Reading Collection Modern Chemistry Chemical Bonding Section Quiz Answers
  10. Overcoming Reading Challenges
    - Dealing with Digital Eye Strain
    - Minimizing Distractions
    - Managing Screen Time
  11. Cultivating a Reading Routine Modern Chemistry Chemical Bonding Section Quiz Answers
    - Setting Reading Goals Modern Chemistry Chemical Bonding Section Quiz Answers
    - Carving Out Dedicated Reading Time
  12. Sourcing Reliable Information of Modern Chemistry Chemical Bonding Section Quiz Answers



- Fact-Checking eBook Content of Modern Chemistry Chemical Bonding Section Quiz Answers
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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