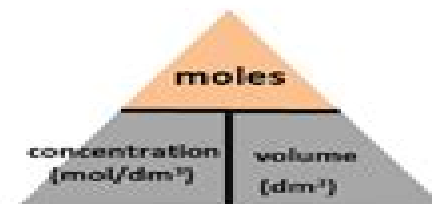


Concentration worksheet

Concentration = Moles / Volume

Moles = Mass / Relative Atomic Mass

$1 \text{ dm}^3 = 1000 \text{ cm}^3$ $0.5 \text{ dm}^3 = 500 \text{ cm}^3$



1) What is the concentration of a solution made when 52 grams of potassium sulfate (K_2SO_4) and dissolved to a volume of 4100 mL?

(Find the moles first, then convert the mL into dm^3 and calculate concentration)

2) 6.80 g of sodium chloride (NaCl) is added to 2750 mL of water. Find the concentration of the solution.

3) Find the volume of a 0.75 mol/dm^3 solution if it contains 39 grams of potassium hydroxide (KOH). (Find moles first)

4) How many grams of hydrochloric acid (HCl) are present in 3.0L of a 0.750 mol/dm^3 solution?

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Harold Clark Metcalfe



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