

Chemistry: Molality

Directions: Using these formulas, solve the following problems.

$$\text{molality} = \text{mol solute} / \text{kg solvent}$$

$$1 \text{ amu} = 1 \text{ gram} / \text{mol}$$

$$\Delta T_{\text{freezing}} = (- \text{number of ions}) (\text{molality}) (k_f)$$

$$\Delta T_{\text{boiling}} = (\text{number of ions}) (\text{molality}) (k_b)$$

k_b and k_f are constants for each individual solvent.

Key



Problem 1: What is the molality of a solution containing 1.23 g of $\text{Hg}(\text{CN})_2$ dissolved in 1.37 g of methanol? Answer: **3.55 m**

Problem 2: What is the molality of a solution containing 16.8 g of NaCl dissolved in 356 g of water? Answer: **0.806 m**

Problem 3: How many grams of NaOH would be required to produce a 3.57 m solution with 473 g of water? Answer: **67.6 g**

Problem 4: How many grams of KCl would you need to dissolve in 114 g of vinegar in order to prepare a 1.96 m solution? Answer: **16.6 g**

Problem 5: Calculate the molality of 5.97 g NaCl dissolved in 471 g of water.
Answer: **0.217 m**

Problem 6: Calculate the molality of 8.91 g Na_2CO_3 dissolved in 595 g of water.
Answer: **0.341 m**

Problem 7: Calculate the molality of 8.44 g KBr dissolved in 112 g of water.
Answer: **0.633 m**

Problem 8: What are the new boiling and freezing points for a solution with 581 g KBr dissolved in 845 grams of water? (k_b water = 0.512°C/m and k_f water = 1.86°C/m)
Answer: new freezing point = **-1.34°C** new boiling point = **106°C**

Problem 9: What are the new boiling and freezing points for a solution with 291 g NaOH dissolved in 1029 grams of water? (k_b water = 0.512°C/m and k_f water = 1.86°C/m)
Answer: new freezing point = **-36.3°C** new boiling point = **107°C**

Molality Answers Key

Lorraine J. Doucette

A red circular graphic with a gradient, appearing as a stylized arrow or a partial circle, located to the right of the author's name.

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