

# The Molar Volume of a Gas

In this experiment, you will determine the molar volume of a gas by conducting a chemical reaction that produces a gas, as shown in the reaction equation below.



You will react a known mass of solid magnesium with an excess of hydrochloric acid, in a sealed vessel, and use the pressure change to calculate molar volume at STP.

## OBJECTIVES

In this experiment, you will

- Measure the gas production of a chemical reaction by a pressure change.
- Determine the molar volume of the gas produced in the reaction.
- Calculate the molar volume of a gas at STP.



Figure 1

## MATERIALS

Vernier computer interface  
computer  
Vernier Gas Pressure Sensor  
Temperature Probe  
600 mL or one liter beaker  
10 mL graduated cylinder

magnesium ribbon, pre-cut  
3.0 M HCl solution  
small beaker for HCl solution  
ruler  
125 mL Erlenmeyer flask  
syringe

# Molar Volume Of Hydrogen Gas Lab Answers

**Yan Bai**



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