

MOLARITY BY DILUTION

79+52 = 132 Name _____

Acids are usually acquired from chemical supply houses in concentrated form. These acids are diluted to the desired concentration by adding water. Since moles of acid before dilution = moles of acid after dilution, and moles of acid = $M \times V$ then, $M_1 \times V_1 = M_2 \times V_2$. Solve the following problems.

1. How much concentrated 18 M sulfuric acid is needed to prepare 250 mL of a 6.0 M solution?

$$V_1 = \frac{M_2 V_2}{M_1} = \frac{(6.0 \text{ M})(250 \text{ mL})}{18 \text{ M}}$$

83 mL

2. How much concentrated 12 M hydrochloric acid is needed to prepare 100 mL of a 2.0 M solution?

$$V_1 = \frac{M_2 V_2}{M_1} = \frac{(2.0 \text{ M})(100 \text{ mL})}{12 \text{ M}}$$

17 mL

3. To what volume should 25 mL of 15 M nitric acid be diluted to prepare a 3.0 M solution?

$$V_2 = \frac{M_1 V_1}{M_2} = \frac{(15 \text{ M})(25 \text{ mL})}{3.0 \text{ M}}$$

125 mL

4. To how much water should 50 mL of 12 M hydrochloric acid be added to produce a 4.0 M solution?

$$V_2 = \frac{M_1 V_1}{M_2} = \frac{(12 \text{ M})(50 \text{ mL})}{4.0 \text{ M}} = 150 \text{ mL} - 50 \text{ mL}$$

100 mL H₂O

5. To how much water should 100 mL of 18 M sulfuric acid be added to prepare a 1.5 M solution?

$$V_2 = \frac{M_1 V_1}{M_2} = \frac{(18 \text{ M})(100 \text{ mL})}{1.5 \text{ M}} = 1200 \text{ mL} - 100 \text{ mL} = 1100 \text{ mL}$$

1.1 L

Molarity Chemistry If8766 Answer Key

JS Bruner



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Table of Contents Molarity Chemistry If8766 Answer Key

1. Understanding the eBook Molarity Chemistry If8766 Answer Key
 - The Rise of Digital Reading Molarity Chemistry If8766 Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Molarity Chemistry If8766 Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Molarity Chemistry If8766 Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Molarity Chemistry If8766 Answer Key
 - Personalized Recommendations
 - Molarity Chemistry If8766 Answer Key User Reviews and Ratings
 - Molarity Chemistry If8766 Answer Key and Bestseller Lists

5. Accessing Molarity Chemistry If8766 Answer Key Free and Paid eBooks
 - Molarity Chemistry If8766 Answer Key Public Domain eBooks
 - Molarity Chemistry If8766 Answer Key eBook Subscription Services
 - Molarity Chemistry If8766 Answer Key Budget-Friendly Options
6. Navigating Molarity Chemistry If8766 Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Molarity Chemistry If8766 Answer Key Compatibility with Devices
 - Molarity Chemistry If8766 Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Molarity Chemistry If8766 Answer Key
 - Highlighting and Note-Taking Molarity Chemistry If8766 Answer Key
 - Interactive Elements Molarity Chemistry If8766 Answer Key
8. Staying Engaged with Molarity Chemistry If8766 Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Molarity Chemistry If8766 Answer Key
9. Balancing eBooks and Physical Books Molarity Chemistry If8766 Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Molarity Chemistry If8766 Answer Key
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Molarity Chemistry If8766 Answer Key
 - Setting Reading Goals Molarity Chemistry If8766 Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Molarity Chemistry If8766 Answer Key
 - Fact-Checking eBook Content of Molarity Chemistry If8766 Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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