

Molarity, Dilutions, and Preparing Solutions

A. Preparing a Stock Solution of Known Concentration

1. Mass of 100-mL volumetric flask, g 59.45
2. Mass of 100-mL volumetric flask and CuSO_4 , g 64.08
3. Mass of CuSO_4 , g 4.63 g
4. Volume of stock solution, mL 100 mL
5. Molarity of stock solution "S," M 0.0290 M
(Show calculations.)
$$\frac{4.63}{159.6} = 0.0290$$

B. Preparing the Spectrophotometer for use

1. Percent transmission of DI H_2O , %T 100 %
2. Absorbance of DI H_2O , A 0

C. Measuring the Absorbance of the Stock Solution

1. Percent transmission of stock solution, %T 43.9
2. Absorbance of stock solution, A 0.357

E. Construction of a Calibration Graph

	"A"	"B"	"C"	"D"
Volume of stock solution, mL	5.00	10.00	15.00	20.00
Volume of prepared solution, mL	50.00	50.00	50.00	50.00
Molarity, M	<u>0.0058</u>			
Percent transmission, %T	<u>105</u>	<u>82.3</u>	<u>79.5</u>	<u>52.6</u>
Absorbance, A	<u>.021</u>	<u>.085</u>	<u>.100</u>	<u>.280</u>

Molarity Concentration Lab Report

Garin Hess

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the name bar.

Molarity Concentration Lab Report:

Laboratory Manual for Principles of General Chemistry J. A. Beran, Mark Lassiter, 2022-08-16 Laboratory Manual for Principles of General Chemistry 11th Edition covers two semesters of a general chemistry laboratory program The material focuses on the lab experiences that reinforce the concepts that not all experimental conclusions are the same and depend on identifying an appropriate experimental procedure selecting the proper apparatus employing the proper techniques systematically analyzing and interpreting the data and minimizing inherent variables As a result of good data a scientific and analytical conclusion is made which may or may not be right but is certainly consistent with the data Experiments write textbooks textbooks don't write experiments A student's scientific literacy grows when experiences and observations associated with the scientific method are encountered Further experimentation provides additional cause effect observations leading to an even better understanding of the experiment The 11th edition's experiments are informative and challenging while offering a solid foundation for technique safety and experimental procedure The reporting and analysis of the data and the pre and post lab questions focus on the intuitiveness of the experiment The experiments may accompany any general chemistry textbook and are compiled at the beginning of each curricular unit An Additional Notes column is included in each experiment's Report Sheet to provide a space for recording observations and data during the experiment Continued emphasis on handling data is supported by the Data Analysis section

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