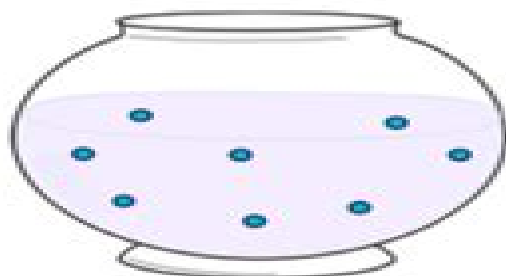


Introduction to Molarity

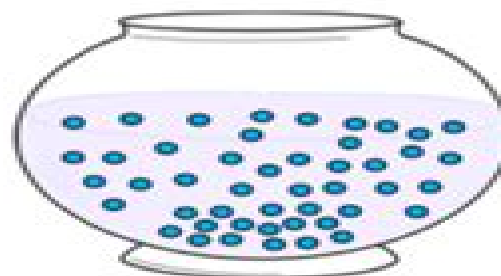
Concentration describes the amount of solute that is present in a certain amount of solution.

- **High concentration** indicates a **large** amount of solute.
- **Low concentration** indicates a **small** amount of solute.

Example: Label the solutions as **high concentration** or **low concentration**.



Low Concentration



High Concentration

Molarity is used to measure the concentration of a solution.

$$\text{Molarity} = \frac{\text{moles of solute}}{\text{liters of solvent}}$$

Example: In every 5 liters of ocean water, there are 20 moles of NaCl (salt).
What is the **molarity** of NaCl in sea water?

$$\text{Molarity} = \frac{\text{moles of solute}}{\text{liters of solvent}}$$

$$\text{Molarity} = \frac{20 \text{ moles NaCl}}{5 \text{ liters water}}$$

$$\text{Molarity} = 4.0 \text{ M or mol/L}$$

Molarity Notes And Worksheet

**Douglas A. Drossman, Ian S.
Grimm, Nicholas J. Shaheen**



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