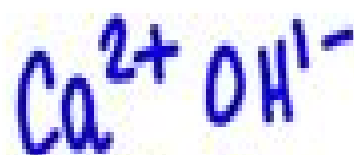


Calculate the molarity of a solution formed by dissolving 24.3 g of calcium hydroxide in 450.0 mL.

$$M = \frac{\text{mol}}{\text{L}} = (24.3 \text{ g}) \left(\frac{\text{mol}}{74.1 \text{ g}} \right) \left(\frac{1}{\text{L}} \right)$$



$$\begin{array}{l} \text{Ca} = 1 (40.1) \\ \text{O} = 2 (16.0) \\ \text{H} = 2 (1.0) \end{array}$$

$$\begin{array}{r} 40.1 \\ 32.0 \\ 2.0 \\ \hline 74.1 \text{ g/mol} \end{array}$$

$$450.0 \text{ mL} \rightarrow .4500 \text{ L}$$

$$\frac{1000 \text{ mL}}{\text{L}}$$

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