

Molarity Calculations

Name _____	Section _____	Date _____
Molarity Calculations		
Directions: Solve each problem for the indicated quantity.		
1. What is the molarity of 1.32 L of a solution of 0.3087 moles of KMnO_4 ?		
2. How many moles of NaOH are present in 2.36 L of a 1.871 M aqueous solution of NaOH ?		
3. What volume in Liters is an aqueous solution of HCl if it has a molarity of 0.850 M and contains 0.580 moles of HCl ?		
4. What is the molarity of 0.1000 moles of KI in 31.7 mL of aqueous solution?		
5. What is the volume in mL of 1.5000 moles of NaClO with a molarity of 0.6341 M?		
6. What is the molarity of an aqueous solution of 2.35 grams of NaCl in a solution with a volume of 1.914 L?		
7. What is the volume of an aqueous solution in mL if it has a molarity of 0.112 M and contains 49.08 grams of dissolved KBr ?		
8. How many grams of H_2CO_3 are dissolved in 2.32 L of an aqueous solution with a molarity of 0.7542 M?		
9. Find the volume of an aqueous solution of KNO_3 in L if it has a molarity of 1.89 M and contains 16.27 grams of KNO_3 . If the solution contains L of water, what is the volume of the dissolved KNO_3 ?		
10. Potassium sulfate has a density of 2.66 $\frac{\text{g}}{\text{mL}}$. What is the molarity of a solution of 7.337 grams of K_2SO_4 and 1.375 L of water?		

$$\text{Molarity} = \frac{\text{Moles of Solute}}{\text{Liters of Solution}}$$

$$M = \frac{\text{mol}}{L}$$

A CHEMISTRY WORKSHEET

10 PROBLEMS

WORKED ANSWER KEY INCLUDED

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Doctor Chemistry



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used *Teaching the Molarity Principle to High School Students* Wilfred Francis Borden, 1974 **Chemistry Lesson**

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