

STOICHIOMETRY AND MOLE CONCEPT

Stoichiometry

Stoichiometry is a section of chemistry that involves using relationships between reactants and/or products in a chemical reaction to determine desired quantitative data and design chemical formula in according to its combining powers.

Writing chemical formula (molecular formula)

The formula of a compound tells us the number of each type of atom in a molecule. The formula of sulphuric acid is H_2SO_4 . This tells us that there are two moles of hydrogen, one mole of sulphur and four moles of oxygen. For ionic compounds, the formula can be worked out by making sure the total charge equals zero. The valence can also be used to determine a chemical formula. The valence represents the total numbers of covalent bonds an atom can form, or the charge on an ion. Valencies can be worked out by using the periodic table:

Naming Covalent Compounds

Many of these compounds have common names such as "bromine", "ammonia" and "carbon dioxide". However, simple covalent compounds are generally named by using **prefixes** to indicate how many atoms of each element are shown in the formula. Also, the ending of the last (most negative) element is changed to -ide.

The prefixes used are **mono-**, **di-**, **tri-**, **tetra-**, **penta-**, **hexa-**, and so forth. The **mono-** prefix is usually not used for the first element in the formula. The "o" and "a" endings of these prefixes commonly are dropped when they are attached to "oxide".

E.g. Carbon dioxide – CO_2

Sulphur trioxide – SO_3

Phosphorous pentachloride PCl_5

Carbon tetrachloride CCl_4

Number of bonds	Prefix
1	mono-
2	di-
3	tri-
4	tetra-
5	penta-
6	hexa-

Naming ionic compounds

Mole Concept And Stoichiometry Fy Bsc Chapter

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Mole Concept And Stoichiometry Fy Bsc Chapter:

Teaching the Mole Concept in Stoichiometry Timothy Scot Albrecht, California State Polytechnic University, Pomona. College of Education & Integrative Studies, 2009 **Formulation and Stoichiometry** Emil J. Margolis, 2012-12-06 The purpose of this book is to interpret more sensitively some of the offerings of the standard text book of general chemistry As a supplement thereto it covers various aspects of formulation and stoichiometry that are frequently treated far too perfunctorily or in many instances are not considered at all The inadequate attention often accorded by the comprehensive text to many topics within its proper purview arises understandably enough from the numerous broad and highly varied objectives set for the first year of the curriculum for modern chemistry in colleges and universities For the serious student this means more often than not the frustrations of questions unanswered The amplification that this book proffers in the immediate area of its subject covers the equations representing internal redox reactions not only of the simple but also of the multiple disproportionations of which the complexities often discourage an undertaking despite the challenge they offer distinctions to be observed in the balancing of equations in contrasting alkali basic and ammonia basic reaction media quantitative contributions made by the ionization or dissociation effects of electrolytes to the colligative properties of their solutions intensive application of the universal reaction principle of chemical equivalence to the stoichiometry of oxidation and reduction *The Mole Concept* Paul Gilchrist, 1982* **The Mole Concept**, 1983 SUMMARY A series of 6 programmes explaining the reasoning behind the development and acceptance of the mole concept *Stoichiometry* Sydney Young, 1918 This introduction was originally prefixed to Dr Findlay's Phase rule which was the first volume of the series issued It belongs properly however to this volume and is therefore included here Note p vii Introduction issued also separately in 1904 **The Teaching of the Mole Concept** John Slater Allen, 1976 *The Mole Concept* Alan Keith Griffiths, 1979 **The Mole Concept : Investigation of an Hierarchical Model** Alan K. (Alan Keith) Griffiths, 1979

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