

Mole Lab

Introduction to The Mole Concept



Introduction

Although technically not a laboratory experiment, this activity certainly helps to drive home the main idea behind the mole concept—that chemists can count out infinitesimally small particles by weighing.

Concepts

- Avogadro's number
- Chemical formulas
- Molar mass or molecular weight

Materials (for each lab station, two or three lab pairs can work at each station)

- Aluminum foil
- Ammonium sulfate, $(\text{NH}_4)_2\text{SO}_4$, 100 g
- Sodium chloride, NaCl , 500 g
- Water, distilled, 500 mL
- Balance, electronic, 0.01-g precision (if only triple beam balances are available, it is best if there is one for each lab pair)
- Beakers, 600-mL, or cut off 2-L soda bottles, 3 labeled
- Cups, blue, clear, and red plastic, 2–3 of each
- Pipets, Beral-type, 2–3
- Spoons, plastic, 4–6

Materials (for the instructor's station, preferably near the center of the room)

- Copper sheet, 12 pieces*
 - Balance, electronic, 0.01-g precision
 - Cups, blue, clear, and red plastic, 1 each
 - Scoring Standard Table
- *See the Preparation section below for sizes of copper pieces.*

Safety Precautions

Ammonium sulfate is slightly toxic by ingestion. Wear chemical splash goggles, chemical-resistant gloves, and a chemical-resistant apron. Wash hands thoroughly with soap and water before leaving the laboratory. Follow all laboratory safety guidelines. Please review current Material Safety Data Sheets for additional safety, handling, and disposal information.

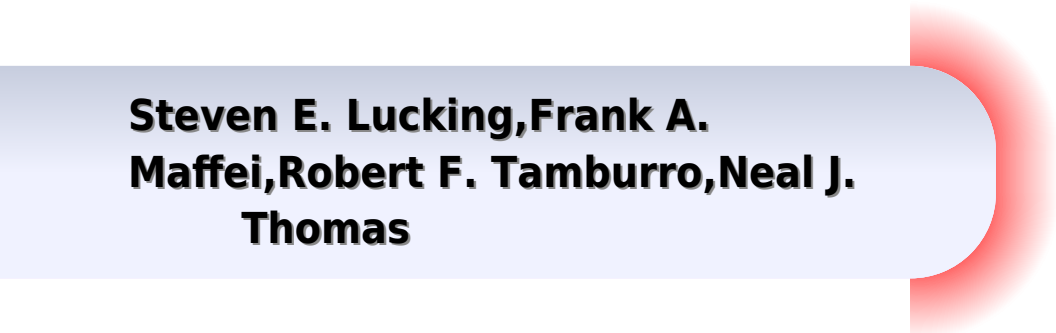
Preparation

1. Label three 600-mL beakers "NaCl," "water," and "ammonium sulfate" and add to each the corresponding substance. *Note:* Do not write the chemical formula for ammonium sulfate on the beaker; students must determine this.
2. Place 2–3 spoons in each beaker containing a solid substance.
3. Place 2–3 pipets in the beaker of water.
4. Cut twelve pieces of copper sheet according to Table 1 and label with the corresponding numbers 1–12. *Note:* The copper pieces can have any weights you want them to. Using the weights given in the table allows you to use the Scoring Standard table as is. If you use other weights, just modify column C in the table accordingly, as shown in Figure 1 on page 2.

Cu Piece	Mass (g)
#1	4.22
#2	4.78
#3	4.98
#4	5.13
#5	5.44
#6	5.97
#7	6.09
#8	6.55
#9	6.79
#10	7.13
#11	7.65
#12	7.99

Mole Lab Student Guide Answer Key

**Steven E. Lucking, Frank A.
Maffei, Robert F. Tamburro, Neal J.
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Resources in Education ,1980 *Pediatric Critical Care Study Guide* Steven E. Lucking, Frank A. Maffei, Robert F. Tamburro, Neal J. Thomas, 2012-04-12 This is the first comprehensive study guide covering all aspects of pediatric critical care medicine It fills a void that exists in learning resources currently available to pediatric critical care practitioners The major textbooks are excellent references but do not allow concise reading on specific topics and are not intended to act as both text and study guide There are also several handbooks available but these are usually written for general pediatric residents and lack the advanced physiology and pathophysiology required for the higher level pediatric critical care practitioner Study Guide for Chemical Principles Thomas Elliott Taylor, 1979 **The Extraordinary Chemistry of Ordinary Things, Study Guide** Carl H. Snyder, David F. Dever, 1995-01-19 Examines the chemistry of the substances of our everyday world Our daily lives are immersed in chemicals an effective way to teach and learn chemistry is by examining the goods and substances that we use in our daily lives and that affect us and our environment **Resources for Teaching Middle School Science** Smithsonian Institution, National Academy of Engineering, National Science Resources Center of the National Academy of Sciences, Institute of Medicine, 1998-04-30 With age appropriate inquiry centered curriculum materials and sound teaching practices middle school science can capture the interest and energy of adolescent students and expand their understanding of the world around them Resources for Teaching Middle School Science developed by the National Science Resources Center NSRC is a valuable tool for identifying and selecting effective science curriculum materials that will engage students in grades 6 through 8 The volume describes more than 400 curriculum titles that are aligned with the National Science Education Standards This completely new guide follows on the success of Resources for Teaching Elementary School Science the first in the NSRC series of annotated guides to hands on inquiry centered curriculum materials and other resources for science teachers The curriculum materials in the new guide are grouped in five chapters by scientific area Physical Science Life Science Environmental Science Earth and Space Science and Multidisciplinary and Applied Science They are also grouped by type core materials supplementary units and science activity books Each annotation of curriculum material includes a recommended grade level a description of the activities involved and of what students can be expected to learn a list of accompanying materials a reading level and ordering information The curriculum materials included in this book were selected by panels of teachers and scientists using evaluation criteria developed for the guide The criteria reflect and incorporate goals and principles of the National Science Education Standards The annotations designate the specific content standards on which these curriculum pieces focus In addition to the curriculum chapters the guide contains six chapters of diverse resources that are directly relevant to middle school science Among these is a chapter on educational software and multimedia programs chapters on books about science and teaching directories and guides to science trade books and periodicals for teachers and students Another section features institutional resources One chapter

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