

MOLE RATIOS IN CHEMICAL EQUATIONS

When the equation is balanced, you can use the **coefficients** to get the **mole ratios** between all substances present to get the **amount** (moles) of each (pure) substance.

Remember, that one mole of a substance is equal to Avogadro's number (N_A) = 6.022×10^{23} particles/mol of **particles** or **entities** of that substance (which could be a reactant or product). Below is a list of what an 'entity' represents for each type of pure substance.

Element – atom e.g. Al

Covalent/Molecular Compound – molecule e.g. H_2SO_4

Ionic Compound – formula unit e.g. $Al_2(SO_4)_3$

There is more than one way to set up/look at a mole ratio, but ultimately the final answer should be the same. There are short cuts too, but you should only use these after you understand what is happening.



MOLE RATIOS



so, if you have 5.00 moles of Al, then how many moles of H_2SO_4 are required to react with the Al?

"cross multiplication" method

$$\begin{array}{l} 2 \text{ Al} : 3 \text{ H}_2\text{SO}_4 \\ \frac{2 \text{ Al}}{5.00 \text{ Al}} = \frac{3 \text{ H}_2\text{SO}_4}{x \text{ mol H}_2\text{SO}_4} \end{array}$$

$$= \frac{3 \text{ H}_2\text{SO}_4 \times 5.00 \text{ Al}}{2 \text{ Al}}$$

G x R/G method

$$\begin{array}{l} G \times \frac{R}{G} = 5.00 \text{ Al} \times \frac{3 \text{ H}_2\text{SO}_4}{2 \text{ Al}} \\ \quad \quad \quad \uparrow \\ \quad \quad \quad \text{(coefficients)} \end{array}$$

$$= 7.50 \text{ mols H}_2\text{SO}_4$$

Mole Ratios In Chemical Reactions Pogil Answers

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