

Calculating Molecular Formulas of Compounds

Name _____	Section _____	Date _____
Molecular Formula		
Directions: Calculate the molecular formula of each compound from the data provided.		
1) A compound consists of 1.25 percent hydrogen and 98.75 percent bromine. What is the molecular formula of the compound if the molar mass of the molecular formula is $80 \frac{g}{mol}$?		
2) Analysis shows a compound contains 24.02 grams of carbon and 75.98 grams of fluorine. If the molar mass of the molecular formula of this compound is $100 \frac{g}{mol}$, what is the molecular formula of the compound?		
3) What is the molecular formula of a compound consisting of 83.62% carbon and 14.37% hydrogen? The molar mass of the molecular formula is $84.2 \frac{g}{mol}$.		
4) A compound with a molar mass of $348 \frac{g}{mol}$ consists of 16.12 grams of nitrogen and 185.88 grams of bromine. What is its molecular formula?		
5) A compound consists of 28.19 grams of nitrogen, 6.13 grams of hydrogen, 20.78 grams of phosphorus, and 42.93 grams of oxygen. If the molar mass of the molecular formula of this compound is $149 \frac{g}{mol}$, what is the molecular formula of the compound?		

6) If the molar mass of the molecular formula of a compound is $140 \frac{g}{mol}$, what is the molecular formula of the compound if it consists of 7.427 grams of chromium and 2.573 grams of carbon?
7) The composition of a compound is 9.49 grams of sodium, 10.13 grams of oxygen, and 13.52 grams of sulfur. If the molar mass of the molecular formula of this compound is $158 \frac{g}{mol}$, what is its molecular formula?
8) A compound is composed of 58.88 grams of barium, 13.74 grams of sulfur, and 27.42 grams of oxygen. If the molar mass of the molecular formula of this compound is $233.40 \frac{g}{mol}$, what is the molecular formula of the compound?
9) If the molar mass of the molecular formula of a compound is $162 \frac{g}{mol}$, what is the molecular formula of the compound if it consists of 2.52 grams of calcium and 68 grams of phosphorus?
10) The molar mass of the molecular formula of a compound is $233.408 \frac{g}{mol}$. What is the molecular formula of the compound if it consists of 29.06 grams of aluminum, 60.54 grams of oxygen, and 15.40 grams of carbon?

A CHEMISTRY WORKSHEET
10 PROBLEMS
3 VERSIONS
WORKED ANSWER KEY

Molecular Formulas Pages 215 216

Gerald James Holton, Stephen G. Brush



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