

Moles and Chemical Formulas

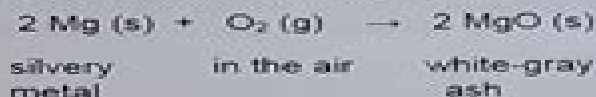
Purpose

- Use mole conversion factors to convert from moles to grams and from grams to moles
- Experimentally determine the simplest formula for magnesium oxide
- Determine the formula of a hydrate

Discussion

The simplest formula of a compound (also called the empirical formula) is the smallest whole number ratio of the atoms. In this experiment, you will conduct two experiments to help you determine the simplest formulas for two different compounds, an oxide of magnesium and a hydrate.

In Part A of this experiment, you will find the number of moles of magnesium ribbon you started with and the number of moles of oxygen that combined with it to form your magnesium oxide product. The reaction is written below.



You can then calculate the simplest formula by finding the ratio of the two compounds by dividing their mole amounts. For example, a different experiment found that 0.20 mole of Al combined with 0.60 mol of Cl. You can find the simplest formula by dividing the larger number by the smaller number. Remember, if your calculator gives you an answer of 3, it is the same as $\frac{3}{1}$.

$$\text{Ratio of moles} = \frac{0.60 \text{ mol Cl}}{0.20 \text{ mol Al}} = \frac{3.0 \text{ mol Cl}}{1.0 \text{ mol Al}}$$

Now you can represent that information as subscripts in the compound's formula.



If the number is a decimal, then both numbers must be multiplied by the same number to obtain the lowest whole number integers, which was discovered by John Dalton and postulated in the Atomic Theory

For example, if you had a Ratio of Moles = $\frac{1.5 \text{ mol O}}{1 \text{ mol N}}$

Both moles could be multiplied by two to give $\frac{3 \text{ mol O}}{2 \text{ mol N}}$

The molar ratio can be re-written as a formula, and would be, N_2O_3

Moles And Chemical Formulas Lab Manual

NARAYAN CHANGDER



Moles And Chemical Formulas Lab Manual:

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