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(a) The symbol equation for the reaction can be represented as shown in Fig. 3.1.

Fig. 3.1

Table 3.1 shows some bond energies.

Bond	H-H	N-N
Bond energy in kJ/mol	435	193

Use the bond energies in Table 3.1 and $\Delta H = -91 \text{ kJ/mol}$ to calculate the bond energy of an N-H bond in kJ/mol.

Use the following steps.

- Calculate the energy needed to break bonds in the reactants.

$$\begin{aligned}
 & (4 \times 435) + 1(193) \\
 & 4 \times 435 + 1(193) \quad \underline{\quad 2150 \quad} \text{ kJ}
 \end{aligned}$$
- Calculate the energy released when bonds form in the products.

$$\begin{aligned}
 & \text{Bond broken} - \text{Bond formed} = -91 \\
 & 2150 - x = -91 \\
 & x = 2150 + 91 \\
 & \quad \quad \quad \underline{\quad 2241 \quad} \text{ kJ}
 \end{aligned}$$
- Calculate the energy of per mole bond.

$$\begin{aligned}
 & 6 \times \text{N-H} = 2241 \\
 & \text{N-H} = \frac{2241}{6} \quad \underline{\quad 373.5 \quad} \text{ kJ/mol}
 \end{aligned}$$

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