

## 7 • Molecular Structure

### STATION 3 • POLARITY

Determine the shape of the molecule and decide whether each compound is polar or nonpolar.

$\text{PH}_3$ $3 \times 1 = 3$ $\text{H} : \underset{\cdot\cdot}{\underset{\cdot\cdot}{\text{P}}} : \text{H}$  Polar	$\text{CF}_4$ $4 \times 4 = 16$ $\begin{array}{c} :\ddot{\text{F}}: \\ :\ddot{\text{F}}: \text{C} :\ddot{\text{F}}: \\ :\ddot{\text{F}}: \end{array}$  non-polar	$\text{BH}_3$ $3 \times 1 = 3$ $\text{H} : \underset{\cdot\cdot}{\underset{\cdot\cdot}{\text{B}}} : \text{H}$  non-polar
$\text{SO}_2$ $6$ $\begin{array}{c} :\ddot{\text{O}}: \\ :\ddot{\text{O}}: \text{S} :\ddot{\text{O}}: \\ :\ddot{\text{O}}: \end{array}$ or $\begin{array}{c} :\ddot{\text{O}}: \\ :\ddot{\text{O}}: \text{S} :\ddot{\text{O}}: \\ :\ddot{\text{O}}: \end{array}$  Polar	$\text{CO}_2$ $12 \times 4 = 16$ $\begin{array}{c} :\ddot{\text{O}}: \\ :\ddot{\text{O}}: \text{C} :\ddot{\text{O}}: \\ :\ddot{\text{O}}: \end{array}$  non-polar	$\text{SCl}_2$ $6 + 6 = 12$ $\begin{array}{c} :\ddot{\text{Cl}}: \\ :\ddot{\text{Cl}}: \text{S} :\ddot{\text{Cl}}: \\ :\ddot{\text{Cl}}: \end{array}$  polar

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### STATION 4 • IMF's AND ATTRACTIONS

Match the statement with the Inter-Particle Forces

A. Dipole-Dipole Interactions

B. Covalent bonds

C. Hydrogen-bonding

D. Ionic Bonds

E. Metallic Bonds

F. London Dispersion Forces

- F 1. Is used to explain why BP of  $\text{Xe} > \text{Kr} > \text{Ar} > \text{Ne} > \text{He}$
- B 2. Is present in ~~Complete~~ all ~~Compounds~~ all ~~Compounds~~
- E 3. Is used to explain why Cu is a good conductor
- D 4. Is used to explain why  $\text{NaCl}(\text{l})$  is a good conductor
- A 5. Is used to explain why  $\text{ICI}$  has a higher BP than  $\text{Br}_2$
- C 6. Is used to explain why boiling point of  $\text{CH}_3\text{OH}$  is higher than  $\text{CH}_4$
- E 7. A lattice of positive ions in a sea of delocalized electrons
- A 8. Positive ends of molecules attract negative ends of other molecules
- E 9. Lattice of alternating positively and negatively charged particles
- C 10. Hydrogen atoms tethered between small, highly electronegative atoms N, O, and F.
- F 11. Polarized electron clouds induce dipoles in their neighboring atoms

# Modeling Chemistry Unit 7 Review Answers

**BM King**



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