

Equilibrium Practice Questions

Name: _____

Equilibrium Questions

1. At an unspecified temperature a 3.00 L flask contains 3.5 mol of HI(g) , 4.1 mol of $\text{H}_2\text{(g)}$ and 0.30 mol of $\text{I}_2\text{(g)}$. Calculate K_c for: $\text{H}_2\text{(g)} + \text{I}_2\text{(g)} \rightleftharpoons 2\text{HI(g)}$
2. At 1200 °C, $K_c = 2.51 \times 10^4$ (no units) for: $\text{H}_2\text{(g)} + \text{I}_2\text{(g)} \rightleftharpoons 2\text{HI(g)}$. What concentration of HI(g) will be produced if 0.250 mol/L of each reactant comes to equilibrium?
3. At 460°C K_c for: $\text{SO}_2\text{(g)} + \text{NO}_2\text{(g)} \rightleftharpoons \text{SO}_3\text{(g)} + \text{NO(g)}$ - is 85.0 (no units). Estimate the product equilibrium concentrations if 0.100 mol/L of each reactant is allowed to come to equilibrium?
4. At 1530 °C, the K_c of: $\text{N}_2\text{(g)} + \text{O}_2\text{(g)} \rightleftharpoons 2\text{NO(g)}$ - is 1.20×10^{-4} (no units). What are the equilibrium concentrations of all the reaction components, after 1.00 mol/L of each of the reactants is allowed to come to equilibrium?

Kinetics And Equilibrium Practice Problems

Ted Lister, Janet Renshaw



Kinetics And Equilibrium Practice Problems:

Survival Guide to General Chemistry Patrick E. McMahon, Rosemary McMahon, Bohdan Khomtchouk, 2019-02-13 This work evolved over thirty combined years of teaching general chemistry to a variety of student demographics The focus is not to recap or review the theoretical concepts well described in the available texts Instead the topics and descriptions in this book make available specific detailed step by step methods and procedures for solving the major types of problems in general chemistry Explanations instructional process sequences solved examples and completely solved practice problems are greatly expanded containing significantly more detail than can usually be devoted to in a comprehensive text Many chapters also provide alternative viewpoints as an aid to understanding Key Features The authors have included every major topic in the first semester of general chemistry and most major topics from the second semester Each is written in a specific and detailed step by step process for problem solving whether mathematical or conceptual Each topic has greatly expanded examples and solved practice problems containing significantly more detail than found in comprehensive texts Includes a chapter designed to eliminate confusion concerning acid base reactions which often persists through working with acid base equilibrium Many chapters provide alternative viewpoints as an aid to understanding This book addresses a very real need for a large number of incoming freshman in STEM fields

New Understanding Chemistry for Advanced Level Third Edition Ted Lister, Janet Renshaw, 2000 Matches the specifications of the Awarding Bodies AQA NEAB AEB OCR and Edexcel This accessible text includes frequent hints questions and examination questions providing support and facilitating study at home It features photographs and comprehensive illustrations with 3D chemical structures

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