Calculating limiting reactants.

Let's look at an example:

$$N_2 + 3H_2 \rightarrow 2NH_3$$



In this reaction, 56 g nitrogen gas (N_2) was added to a chamber containing 14 g hydrogen gas (H_2) . Which was the limiting reactant? $(A_r: H=1, N=14)$

First steps?

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Veterinary and Human Toxicology, 1982 Theoretical Foundations of Chemical Engineering, 1980 Journal of the American Chemical Society American Chemical Society, 1970

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