

Name _____ Date _____ Class _____

CHAPTER 12

STUDY GUIDE FOR CONTENT MASTERY

Section 12.3 Limiting Reactants

In your textbook, read about why reactions stop and how to determine the limiting reactant.

Study the diagram showing a chemical reaction and the chemical equation that represents the reaction. Then complete the table. Show your calculations for questions 15–17 in the space below the table.



The molar masses of O_2 , NO , and NO_2 are 32.00 g/mol, 30.01 g/mol, and 46.01 g/mol, respectively.

Amount of O_2	Amount of NO	Amount of NO_2	Limiting Reactant	Amount and Mass of Excess Reactant
1 molecule	2 molecules	2 molecules	none	none
4 molecules	4 molecules	4 molecules	NO	2 molecules O_2
2 molecules	4 molecules	4 molecules	O_2	4 molecules NO
1.00 mol	2.00 mol	2.00 mol	none	none
4.00 mol	4.00 mol	4.00 mol	NO	2.00 mol O_2
2.00 mol	4.00 mol	4.00 mol	O_2	4.00 mol NO
1.00 mol	4.00 mol	2.00 mol	NO	3.00 mol O_2
0.500 mol	2.00 mol	1.00 mol	NO	1.50 mol O_2
0.200 mol	0.200 mol	0.200 mol	NO	0.400 mol O_2
32.00 g	60.02 g	92.02 g	none	none
64.00 g	60.02 g	46.01 g	O_2	30.12 g NO
64.00 g	30.01 g	28.76 g	O_2	1.24 g NO

balanced equation mole ratio = 2 mol NO /1 mol O_2

$16.00 \text{ g } O_2 \times 1 \text{ mol } O_2 / 32.00 \text{ g } O_2 = 0.5125 \text{ mol } O_2$

$20.00 \text{ g } NO \times 1 \text{ mol } NO / 30.01 \text{ g } NO = 0.6664 \text{ mol } NO$

actual mole ratio = $0.6664 \text{ mol } NO / 0.5125 \text{ mol } O_2 = 1.302 \text{ mol } NO / 1.000 \text{ mol } O_2$

Because the actual mole ratio of NO/O_2 is larger than the balanced equation mole ratio of NO/O_2 , there is an excess of NO ; O_2 is the limiting reactant.

Mass of NO used = $0.5125 \text{ mol } O_2 \times 2 \text{ mol } NO / 1 \text{ mol } O_2 = 0.4250 \text{ mol } NO$

$0.4250 \text{ mol } NO \times 30.01 \text{ g } NO / 1 \text{ mol } NO = 12.76 \text{ g } NO$

Mass of NO_2 produced = $0.4250 \text{ mol } NO \times 2 \text{ mol } NO_2 / 1 \text{ mol } NO = 0.8500 \text{ mol } NO_2$

Excess NO = $20.00 \text{ g } NO - 12.76 \text{ g } NO = 7.24 \text{ g } NO$

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CHAPTER 13

STUDY GUIDE FOR CONTENT MASTERY

Section 13.4 Percent Yield

In your textbook, read about the yields of products.

Study the diagram and the example problem.



Example Problem: The following chemical equation represents the production of palladium, a valuable metal in the manufacturing of some semiconductor devices.



In one experiment, the reaction yielded 7.42 g of the metal from a 7.00-g sample of palladium. Determine the percent yield of this reaction. The molar masses of Ga and Ga_2O_3 are 69.72 g/mol and 157.44 g/mol, respectively.

Use the information in the diagram and example problem to evaluate each value or expression below. If the value or expression is correct, write correct. If it is incorrect, write the correct value or expression.

1. actual yield unknown 7.42 g Ga_2O_3
2. mass of reactant: 7.00 g Ga correct
3. number of moles of reactant: $7.00 \text{ g } Ga \times \frac{1 \text{ mol } Ga}{69.72 \text{ g } Ga} = 0.100 \text{ mol } Ga$
4. number of moles of product: $0.100 \text{ mol } Ga \times \frac{2 \text{ mol } Ga_2O_3}{1 \text{ mol } Ga} = 0.200 \text{ mol } Ga_2O_3$
5. theoretical yield: $0.200 \text{ mol } Ga_2O_3 \times \frac{157.44 \text{ g } Ga_2O_3}{1 \text{ mol } Ga_2O_3} = 31.49 \text{ g } Ga_2O_3$ correct
6. percent yield: $\frac{7.42 \text{ g } Ga_2O_3}{31.49 \text{ g } Ga_2O_3} \times 100 = 23.6\%$

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The novel begins with Richard Papen, a bright but troubled young man, arriving at Hampden College. Richard is immediately drawn to the group of students who call themselves the Classics Club. The club is led by Henry Winter, a brilliant and charismatic young man. Henry is obsessed with Greek mythology and philosophy, and he quickly draws Richard into his world. The other members of the Classics Club are equally as fascinating. Bunny Corcoran is a wealthy and spoiled young man who is always looking for a good time. Charles Tavis is a quiet and reserved young man who is deeply in love with Henry. Camilla Macaulay is a beautiful and intelligent young woman who is drawn to the power and danger of the Classics Club. The students are all deeply in love with Morrow, and they are willing to do anything to please him. Morrow is a complex and mysterious figure, and he seems to be manipulating the students for his own purposes. As the students become more involved with Morrow, they begin to commit increasingly dangerous acts. The Secret History is a masterful and gripping novel that will keep you wondering until the very end. The novel is a cautionary tale about the dangers of obsession and the power of evil.

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