

Properties of Acids and Bases

(Adapted from Flinn Scientific Acid - Base Test Kit I - #AP4567)

Introduction

Battery acid, stomach acid, acid rain - just a few acids in our everyday life! What does it mean when something is acidic? How do acids differ from bases? In this lab, the physical and chemical properties of a variety of acids and bases will be examined.

Concepts

- Acids vs. Bases
- Indicators
- pH

Background

Acids and bases make up two groups of substances that can be categorized by their physical and chemical properties. Let's take a look at the distinguishing properties:

1. **Taste:** Acids, such as lemons or oranges, taste sour. Bases, such as soap, taste bitter. *Note:* Taste should never be used to identify a lab chemical or unknown substance. This lab does not include a test using taste.
2. **Feel:** Acid solutions do not feel much different than water; however, they sting if they contact broken skin. Base solutions have a slippery feeling. *Note:* Lab chemicals or unknown substances should never be touched with the bare skin. This lab does not include a test using feel.
3. **Indicators:** An indicator is a chemical compound, either on a test paper or in a solution, that changes color depending on the acidity or basicity of a solution and, thus, is used to test for the presence of acids or bases. There are many different kinds of indicators. Litmus paper and phenolphthalein indicator solution are two of the most common.
 - Blue litmus paper turns red when dipped in an acid solution of pH 1-6.
 - Red litmus paper turns blue when dipped in a basic solution of pH 8-14.
 - In a neutral solution with pH 7, each paper retains its original color.
 - Phenolphthalein indicator solution turns a pink color in a basic solution but remains colorless (sometimes cloudy) in an acidic or neutral solution.
4. **Reaction with carbonates:** While most bases do not react with carbonates, acids react with carbonates to release carbon dioxide gas. For example:
$$\text{BaCO}_3 + 2\text{HCl} \rightarrow \text{BaCl}_2 + \text{CO}_2 + \text{H}_2\text{O}$$
5. **Reaction with metals:** While most bases do not react with metals, acids react with active metals to release hydrogen gas. For example:
$$\text{Fe} + 2\text{HCl} \rightarrow \text{FeCl}_2 + \text{H}_2$$

Lab 20 Properties Of Acids And Bases

Mary Ethel Jones



Lab 20 Properties Of Acids And Bases:

Unveiling the Power of Verbal Artistry: An Psychological Sojourn through **Lab 20 Properties Of Acids And Bases**

In a global inundated with screens and the cacophony of quick transmission, the profound power and mental resonance of verbal beauty usually diminish in to obscurity, eclipsed by the continuous onslaught of noise and distractions. However, set within the musical pages of **Lab 20 Properties Of Acids And Bases**, a charming work of literary elegance that pulses with organic emotions, lies an memorable trip waiting to be embarked upon. Published with a virtuoso wordsmith, this enchanting opus manuals visitors on a mental odyssey, delicately exposing the latent potential and profound affect embedded within the complicated internet of language. Within the heart-wrenching expanse with this evocative examination, we shall embark upon an introspective exploration of the book is central subjects, dissect its charming publishing type, and immerse ourselves in the indelible impact it leaves upon the depths of readers souls.

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