Names	Class:	Date:	ID: A

Chapter 10 test

Matching

Match each item with the correct statement below.

- a. representative particle
 d. percent composition
- mole
 standard temperature and pressure
- c. Avogadro's number f. empirical formula
- the number of representative particles of a substance present in 1 mole of that substance
 - 2. an atom, an ion, or a molecule, depending upon the way a substance commonly exists
 - 3. the SI unit used to measure amount of substance
 - 4. 0°C and 1 atm
 - 5. the percent by mass of each element in a compound
- 6. the smallest whole number ratio of the atoms in a compound

Multiple Choice

Identify the letter of the choice that best completes the statement or answers the question.

- 7. What SI unit is used to measure the number of representative particles in a substance?
 - a, kilogram

c. kelvin

b. ampere

- d. mole
- How many hydrogen atoms are in 5 molecules of isopropyl alcohol, C₃H₂O?
 - a. $5 \times (6.02 \times 10^{23})$

c. 35

ъ. 5

- d. 35 × (6.02 × 10²³)
- All of the following are equal to Avogadro's number EXCEPT _____
 - a. the number of atoms of bromine in 1 mol Br₂
 - b. the number of atoms of gold in 1 mol Au
 - c. the number of molecules of nitrogen in 1 mol N₂
 - d. the number of molecules of carbon monoxide in 1 mol CO
- ____ 10. How many moles of tungsten atoms are in 4.8 × 10²⁹ atoms of tungsten?
 - a. 8.0 × 10² moles

c. 1.3×10^{-1} moles

b. 8.0×10^4 moles

- d. 1.3×10^{-2} moles
- _____ 11. How many atoms are in 0.075 mol of titanium?
 - a. 1.2×10^{-25}

c. 6.4 × 10²

b. 2.2×10^{24}

- d. 4.5×10^{23}
- 12. How many molecules are in 2.10 mol CO₂?
 - a. 2.53 × 10²⁴ molecules

c. 3.49×10^{-24} molecules

b. 3.79 × 10²⁴ molecules

d. 1.26 × 10²⁴ molecules

Modern Chemistry Chapter 10 Test

K Morrison

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