

Chemistry Chapter 5 Exam

Multiple Choice

Identify the letter of the choice that best completes the statement or answers the question.

- _____ 1. The idea of arranging the elements in the periodic table according to their chemical and physical properties is attributed to
 - a. Mendeleev.
 - b. Moseley.
 - c. Bohr.
 - d. Ramsay.
- _____ 2. Mendeleev left spaces in his periodic table and predicted several elements and their
 - a. atomic numbers.
 - b. colors.
 - c. properties.
 - d. radioactivity.
- _____ 3. Mendeleev attempted to organize the chemical elements based on their
 - a. symbols.
 - b. properties.
 - c. atomic numbers.
 - d. electron configurations.
- _____ 4. Mendeleev noticed that properties of elements usually repeated at regular intervals when the elements were arranged in order of increasing
 - a. atomic number.
 - b. density.
 - c. reactivity.
 - d. atomic mass.
- _____ 5. Mendeleev is credited with developing the first successful
 - a. periodic table.
 - b. method for determining atomic number.
 - c. test for radioactivity.
 - d. use of X rays.
- _____ 6. Mendeleev did not always list elements in his periodic table in order of increasing atomic mass because he grouped together elements with similar
 - a. properties.
 - b. atomic numbers.
 - c. densities.
 - d. colors.
- _____ 7. In developing his periodic table, Mendeleev listed on cards each element's name, atomic mass, and
 - a. atomic number.
 - b. electron configuration.
 - c. isotopes.
 - d. properties.
- _____ 8. Mendeleev predicted that the spaces in his periodic table represented
 - a. isotopes.
 - b. radioactive elements.
 - c. permanent gaps.
 - d. undiscovered elements.
- _____ 9. Mendeleev's table was called periodic because the properties of the elements
 - a. showed no pattern.
 - b. occurred at repeated intervals called periods.
 - c. occurred at regular time intervals called periods.
 - d. were identical.
- _____ 10. The person whose work led to a periodic table based on increasing atomic number was
 - a. Moseley.
 - b. Mendeleev.
 - c. Rutherford.
 - d. Cannizzaro.
- _____ 11. Moseley's work led to the realization that elements with similar properties occurred at regular intervals when the elements were arranged in order of increasing
 - a. atomic mass.
 - b. density.
 - c. radioactivity.
 - d. atomic number.
- _____ 12. Who used his experimental evidence to determine the order of the elements according to atomic number?
 - a. Meyer
 - c. Stas

Modern Chemistry Chapter 5 2 Review The Answer Key

J Spring



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Medicinal Chemistry Norma K Dunlap, Donna M Huryn, 2025-11-28 This book provides a concise introduction to the medicinal chemistry of therapeutically active compounds Written in an accessible style it offers a brief overview of the drug development process from discovery to approved drug Emphasizing medicinal chemistry strategies the content is organized by target as well as therapeutic area drug classes It covers existing drugs and the discovery process that led to their development Aimed at students and professionals interested in drug discovery Medicinal Chemistry requires little prior knowledge making it ideal for classroom and training seminar use Written by experts in the field this outstanding new edition explains medicinal chemistry on chemical mechanism terms allowing for a more interesting and user friendly text Includes valuable insights toward the various pathways taken by the pharmaceutical industry in drug discoveries Enhanced text includes questions raised and suggestions made from students in the authors medicinal chemistry class Content provided to enhance the text includes a solutions manual PowerPoint slides for lecture outlines and especially for the earlier chapters a selection of videos An accompanying website will be updated to include additional material 1 Journal Club exercises current journal articles with guiding questions for student analysis two per chapter 2 additional case studies 3 organic synthesis analysis selected syntheses both discovery and process will be discussed and 4 updates on current advances in the field of medicinal chemistry The book benefits a diverse audience including upper level undergraduates and graduates who are studying in the fields of medicinal chemistry and drug discovery as well as scientists working in the pharmaceutical industry

Modern Chemistry Holt Rinehart & Winston, Holt, Rinehart and Winston Staff, 2001 **Chemical Fundamentals of Geology and Environmental Geoscience** Robin Gill, 2015-01-27 Chemical principles are fundamental to the Earth sciences and geoscience students increasingly require a firm grasp of basic chemistry to succeed in their studies The enlarged third edition of this highly regarded textbook introduces the student to such geo relevant chemistry presented in the same lucid and accessible style as earlier editions but the new edition has been strengthened in its coverage of environmental geoscience and incorporates a new chapter introducing isotope geochemistry The book comprises three broad sections The first Chapters 1 4 deals with the basic physical chemistry of geological processes The second Chapters 5 8 introduces the wave mechanical view of the atom and explains the various types of chemical bonding that give Earth materials their diverse and distinctive properties The final chapters 9 11 survey the geologically relevant elements and isotopes and explain their formation and their abundances in the cosmos and the Earth The book concludes with an extensive glossary of terms appendices cover basic maths explain basic solution chemistry and list the chemical elements and the symbols units and constants used in the book **Merrill Chemistry** Robert C. Smoot, Smoot, Richard G. Smith, Jack Price, 1998 **The Dyer, Calico Printer, Bleacher, Finisher & Textile Review**, 1924 **General Chemistry** Ralph H. Petrucci, 1989 *Modern Organic Synthesis in the Laboratory* Jie Jack Li, Chris Limberakis, Derek A. Pflum, 2007-09-10 This collection intends to serve

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