

Q.3 The phenomenon in which an ionic compound shows more than one crystalline form is called:

- A) Polymorphism C) Isomorphism
B) Allotropy D) Isomerism

Q.4 All of the following factors affect vapour pressure of liquid EXCEPT:

- A) Nature of liquid C) Surface area
B) Temperature D) Intermolecular forces

Q.7 Which of the following gases shows more non-ideal behaviour?

- A) O_2 C) N_2
B) CO_2 D) H_2

Q.8 Which of the following equations is used for real gases?

- A) $PV = nRT$ C) $\frac{PV}{T} = \frac{PV_0}{T_0}$
B) $PV = \frac{1}{3} n m C^2$ D) $\left(P + \frac{a}{V^2} \right) (V - nb) = nRT$

Q.9 The gas laws can be summarised in the ideal gas equation $PV = nRT$ which of the following statements is / are incorrect?

- A) One mole of any ideal gas occupies the same volume under the same conditions of temperature and pressure.
B) The density of an ideal gas at constant pressure is inversely proportional to temperature.
C) Volume of a given mass of a gas increases two times if temperature is raised from 27°C to 54°C at constant pressure.
D) Both A and B

Q.10 Which of the following statements does not correctly represent the behaviour of an ideal gas?

- A) $PV = C$ C) $PM = nRT$
B) $P \propto CT$ D) $P \propto \frac{1}{d}$

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