

**CHAPTER 9 REVIEW*****Stoichiometry*****MIXED REVIEW****SHORT ANSWER** Answer the following questions in the space provided.

- 1.**
- Given the following equation:
- $\text{C}_3\text{H}_8(\text{g}) + x\text{O}_2(\text{g}) \rightarrow 3\text{CO}_2(\text{g}) + 2\text{H}_2\text{O}(\text{g})$

- \_\_\_\_\_ a. What is the value of the coefficient  $x$  in this equation?
- \_\_\_\_\_ b. What is the molar mass of  $\text{C}_3\text{H}_8$ ?
- \_\_\_\_\_ c. What is the mole ratio of  $\text{O}_2$  to  $\text{H}_2\text{O}$  in the above equation?
- \_\_\_\_\_ d. How many moles are in an 8.0 g sample of  $\text{C}_3\text{H}_8$ ?
- \_\_\_\_\_ e. If  $z$  mol of  $\text{C}_3\text{H}_8$  react, how many moles of  $\text{CO}_2$  are produced, in terms of  $z$ ?

- 2.**
- a. What is meant by
- ideal conditions*
- relative to stoichiometric calculations?

\_\_\_\_\_

\_\_\_\_\_

- b. What function do ideal stoichiometric calculations serve?

\_\_\_\_\_

\_\_\_\_\_

- c. Are actual yields typically larger or smaller than theoretical yields?

\_\_\_\_\_

**PROBLEMS** Write the answer on the line to the left. Show all your work in the space provided.

- 3.**
- Assume the reaction represented by the following equation goes all the way to completion:



- \_\_\_\_\_ a. If 6 mol of  $\text{H}_2$  are consumed, how many moles of  $\text{NH}_3$  are produced?
- \_\_\_\_\_ b. How many grams are in a sample of  $\text{NH}_3$  that contains  $3.0 \times 10^{23}$  molecules?

# Modern Chemistry Stoichiometry Ch 9 Review Answers

**SJ Ball**



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