

CHEM 101 – Recitation Week 9 - Answer Key

Ch. 9, 38. The air in an inflated balloon (defined as the system) warms over a toaster and absorbs 115 J of heat. As it expands, it does 77 kJ of work. What is the change in internal energy for the system?

Concepts needed: heat (q) vs. work (w)
conservation of energy;
units of measure conversion

Answer:

$$\begin{aligned}\Delta E &= q + w \\ q &= +115 \text{ J} \\ w &= -77 \times 10^3 \text{ J} \\ \Delta E &= (115 \text{ J}) + (-77 \times 10^3 \text{ J}) \\ \Delta E &\approx -77 \text{ KJ}\end{aligned}$$

Ch. 9, 40. A kilogram of aluminum metal and a kilogram of water are each warmed to 75 °C and placed in two identical insulated containers. One hour later, the two containers are opened, and the temperature of each substance is measured. The aluminum has cooled to 35 °C, while the water has cooled only to 66 °C. Explain this difference.

Concepts needed: definition of specific heat: $c = q/m\Delta T$ J/(g °C)
conservation of energy;

Answer:

Since the same amount of heat was given to the same mass of substance, the substance with the higher specific heat (in this case, water) will cool down less.

Mathematically, this can be shown by the law of conservation of energy: let "q" be the heat released

$$\left. \begin{aligned}q &= m c_{\text{H}_2\text{O}} \Delta T_{\text{H}_2\text{O}} \\ q &= m c_{\text{Al}} \Delta T_{\text{Al}}\end{aligned} \right\} \rightarrow \cancel{m} c_{\text{H}_2\text{O}} \Delta T_{\text{H}_2\text{O}} = \cancel{m} c_{\text{Al}} \Delta T_{\text{Al}}$$

$c_{\text{H}_2\text{O}} \Delta T_{\text{H}_2\text{O}} = c_{\text{Al}} \Delta T_{\text{Al}}$

When c increases, ΔT must decrease to keep $c \Delta T$ constant !

Ch. 9, 44. Find the mass of a substance at 23 °C that absorbs 1.95×10^3 J of heat. The final temperature and the specific heat of the substance are indicated below.

c. Ethanol: $T_f = 44.2$ °C $c = 2.42$ J/g·C

Concepts needed: specific heat definition: $c = q/m\Delta T$ J/(g °C)
dimensional analysis

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R Barnett



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