



# Chemistry Formulas



Dilution

$$M_1V_1 = M_2V_2$$

Energy of a photon

$$E_{\text{photon}} = hf = hc/\lambda$$

Pressure

$$P = F/A$$

Average Kinetic Energy

$$KE_{\text{avg}} = \frac{3}{2} RT$$

Ideal Gas Law

$$PV = nRT$$

Combined Gas Law

$$\frac{P_1V_1}{n_1T_1} = \frac{P_2V_2}{n_2T_2}$$

Dalton's Law of Partial Pressures

$$P_{\text{total}} = P_A + P_B + \dots$$

Dalton's Law of Partial Pressures

$$P_A = \chi_A P_{\text{total}}$$

Graham's Law of Effusion

$$\frac{r_1}{r_2} = \sqrt{\frac{M_2}{M_1}}$$

Henry's Law

$$P_A = k_H[A]$$

Freezing Point Depression

$$\Delta T_F = -iK_F M$$

Boiling Point Elevation

$$\Delta T_B = iK_B m$$

Vapor Pressure Depression (Raoult's Law)

$$P_{\text{soln}} = \chi_{\text{soln}} P_{\text{soln}}^0$$

Osmotic Pressure ( $\pi$ )

$$\pi = iMRT$$

Arrhenius Equation

$$k = Ae^{-E_a/RT}$$

Equilibrium Constant Expressions

$$K_c = \frac{[\text{products}]}{[\text{reactants}]}$$

Equilibrium Constant Expressions

$$K_{eq} = \frac{k_{\text{forward}}}{k_{\text{reverse}}}$$

Equilibrium Constant Expressions

$$K_P = \frac{P_{\text{products}}}{P_{\text{reactants}}}$$

Reaction Quotient (Q)

$$Q = \frac{[\text{products}]}{[\text{reactants}]}$$

Solubility Product Constant ( $K_{sp}$ )

$$K_{sp} = \frac{[\text{products}]}{[\text{reactants}]}$$

pH & pOH

$$pH = -\log[H^+]$$

pH & pOH

$$pOH = -\log[OH^-]$$

pH & pOH

$$pH + pOH = 14$$

$pK_a$  &  $pK_b$

$pK_a$  &  $pK_b$

$pK_a$  &  $pK_b$

Neutralization Reaction

Buffers

Enthalpy of Formation

First Law of Thermodynamics

Pressure Volume Work

Calorimetry Thermal Energy (q)

Heat Curves & Thermal Energy (q)

Heat Curves & Thermal Energy (q)

Heat Curves & Thermal Energy (q)

Entropy (S)

Bond Dissociation Energy

Gibb's Free Energy ( $\Delta G$ )

Gibb's Free Energy ( $\Delta G$ )

Gibb's Free Energy ( $\Delta G$ )

Standard Cell Potential

Standard Cell Potential

Nernst Equation

Kinetics (always 1st order)

Kinetics (always 1st order)

Kinetics (always 1st order)

Nuclear Binding Energy

$$pK_a = -\log[K_a]$$

$$pK_b = -\log[K_b]$$

$$pK_a + pK_b = 14$$

$$n_a M_a V_a = n_b M_b V_b$$

$$pH = pK_a + \log[A^-]/[HA]$$

$$\Delta H_f = \sum \Delta H_f^{\circ}(\text{product}) - \sum \Delta H_f^{\circ}(\text{reactants})$$

$$\Delta E = q + w$$

$$w = -P\Delta V$$

$$q = -C_{\text{calorimeter}}\Delta T$$

$$q = mc\Delta T$$

$$q = m\Delta H_{\text{fusion}}$$

$$q = m\Delta H_{\text{vaporization}}$$

$$\Delta S = \ln S_{\text{products}} - \ln S_{\text{reactants}}$$

$$\Delta H = \sum \Delta H_{\text{reactants}} - \sum \Delta H_{\text{products}}$$

$$\Delta G^{\circ} = \Delta H^{\circ} - T\Delta S^{\circ}$$

$$\Delta G = \Delta G^{\circ} + RT\ln Q$$

$$\Delta G^{\circ} = -RT^{\circ}\ln(K_{eq})$$

$$E^{\circ} = E^{\circ}_{\text{reduction}} + E^{\circ}_{\text{oxidation}}$$

$$E^{\circ} = E^{\circ}_{\text{cathode}} + E^{\circ}_{\text{anode}}$$

$$E_{\text{cell}} = E^{\circ} - \frac{0.0592}{n} \log Q$$

$$N = N_0 e^{-kt}$$

$$\ln N = \ln N_0 - kt$$

$$t_{1/2} = \frac{0.693}{k}$$

$$E = \Delta mc^2$$

# Modern Chemistry Ch 12 Tests A And

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1. Understanding the eBook Modern Chemistry Ch 12 Tests A And
  - The Rise of Digital Reading Modern Chemistry Ch 12 Tests A And
  - Advantages of eBooks Over Traditional Books
2. Identifying Modern Chemistry Ch 12 Tests A And
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Modern Chemistry Ch 12 Tests A And
  - User-Friendly Interface
4. Exploring eBook Recommendations from Modern Chemistry Ch 12 Tests A And
  - Personalized Recommendations
  - Modern Chemistry Ch 12 Tests A And User Reviews and Ratings
  - Modern Chemistry Ch 12 Tests A And and Bestseller Lists
5. Accessing Modern Chemistry Ch 12 Tests A And Free and Paid eBooks
  - Modern Chemistry Ch 12 Tests A And Public Domain eBooks
  - Modern Chemistry Ch 12 Tests A And eBook Subscription Services
  - Modern Chemistry Ch 12 Tests A And Budget-Friendly Options
6. Navigating Modern Chemistry Ch 12 Tests A And eBook Formats
  - ePub, PDF, MOBI, and More
  - Modern Chemistry Ch 12 Tests A And Compatibility with Devices
  - Modern Chemistry Ch 12 Tests A And Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Modern Chemistry Ch 12 Tests A And
  - Highlighting and Note-Taking Modern Chemistry Ch 12 Tests A And
  - Interactive Elements Modern Chemistry Ch 12 Tests A And

8. Staying Engaged with Modern Chemistry Ch 12 Tests A And
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Modern Chemistry Ch 12 Tests A And
9. Balancing eBooks and Physical Books Modern Chemistry Ch 12 Tests A And
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Modern Chemistry Ch 12 Tests A And
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Modern Chemistry Ch 12 Tests A And
  - Setting Reading Goals Modern Chemistry Ch 12 Tests A And
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Modern Chemistry Ch 12 Tests A And
  - Fact-Checking eBook Content of Modern Chemistry Ch 12 Tests A And
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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