

Chemistry 12 – Unit I – Reaction Kinetics
Kinetics Review Test

1. Consider the following reaction mechanism:

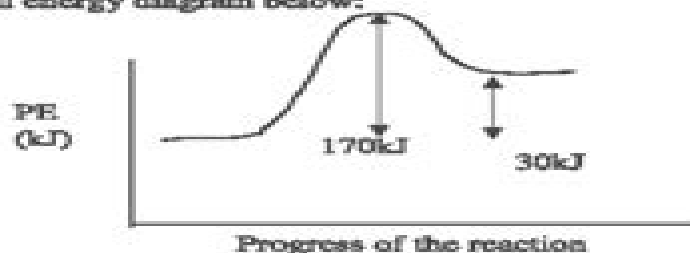


The chemical species MX is a(n) _____.

- A. catalyst B. inhibitor C. product D. reaction intermediate
2. Consider the following reaction:

$$2 \text{N}_2\text{O}_5(g) \rightarrow 4 \text{NO}_2(g) + \text{O}_2(g)$$
 At a certain temperature the rate of decomposition of $\text{N}_2\text{O}_{5(g)}$ is $2.5 \times 10^{-6} \text{ mol/s}$. The rate of formation of NO_2 is _____.
- A. $1.0 \times 10^{-5} \text{ mol/s}$ B. $1.3 \times 10^{-6} \text{ mol/s}$ C. $2.5 \times 10^{-6} \text{ mol/s}$ D. $5.0 \times 10^{-6} \text{ mol/s}$
3. Which of the following factors affect the rates of *both* homogeneous and heterogeneous reactions?
- | | |
|--------------------------------|-----------------------------------|
| I nature of the reactants | II presence of a catalyst |
| III temperature of system | IV concentration of reactants |
- A. I and IV only B. II and III only C. II, III, and IV only D. I, II, III, and IV
4. Which of the following equations represents an endothermic reaction?
- A. $\text{N}_2\text{O}_{5(g)} + 59 \text{ kJ} \rightarrow 2 \text{NO}_{2(g)}$ B. $2 \text{H}_{2(g)} + \text{O}_{2(g)} \rightarrow 2 \text{H}_2\text{O}_{(g)} + 572 \text{ kJ}$
 C. $2 \text{BrCl}_{(g)} - 29.3 \text{ kJ} \rightarrow \text{Br}_{2(g)} + \text{Cl}_{2(g)}$ D. $2 \text{H}_{2(g)} + \text{O}_{2(g)} \rightarrow 2 \text{H}_2\text{O}_{(g)} ; \Delta H = -572 \text{ kJ}$

5. Consider the potential energy diagram below:



The activation energy for the reverse reaction is _____.

- A. 30 kJ B. 140 kJ C. 170 kJ D. 200 kJ

Modern Chemistry Ch 17 Reaction Kinetics Test

Daniela Niemeyer



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Table of Contents Modern Chemistry Ch 17 Reaction Kinetics Test

1. Understanding the eBook Modern Chemistry Ch 17 Reaction Kinetics Test
 - The Rise of Digital Reading Modern Chemistry Ch 17 Reaction Kinetics Test
 - Advantages of eBooks Over Traditional Books
2. Identifying Modern Chemistry Ch 17 Reaction Kinetics Test
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Modern Chemistry Ch 17 Reaction Kinetics Test
 - User-Friendly Interface
4. Exploring eBook Recommendations from Modern Chemistry Ch 17 Reaction Kinetics Test
 - Personalized Recommendations
 - Modern Chemistry Ch 17 Reaction Kinetics Test User Reviews and Ratings
 - Modern Chemistry Ch 17 Reaction Kinetics Test and Bestseller Lists
5. Accessing Modern Chemistry Ch 17 Reaction Kinetics Test Free and Paid eBooks
 - Modern Chemistry Ch 17 Reaction Kinetics Test Public Domain eBooks
 - Modern Chemistry Ch 17 Reaction Kinetics Test eBook Subscription Services
 - Modern Chemistry Ch 17 Reaction Kinetics Test Budget-Friendly Options
6. Navigating Modern Chemistry Ch 17 Reaction Kinetics Test eBook Formats

- ePub, PDF, MOBI, and More
 - Modern Chemistry Ch 17 Reaction Kinetics Test Compatibility with Devices
 - Modern Chemistry Ch 17 Reaction Kinetics Test Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Modern Chemistry Ch 17 Reaction Kinetics Test
 - Highlighting and Note-Taking Modern Chemistry Ch 17 Reaction Kinetics Test
 - Interactive Elements Modern Chemistry Ch 17 Reaction Kinetics Test
 8. Staying Engaged with Modern Chemistry Ch 17 Reaction Kinetics Test
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Modern Chemistry Ch 17 Reaction Kinetics Test
 9. Balancing eBooks and Physical Books Modern Chemistry Ch 17 Reaction Kinetics Test
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Modern Chemistry Ch 17 Reaction Kinetics Test
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Modern Chemistry Ch 17 Reaction Kinetics Test
 - Setting Reading Goals Modern Chemistry Ch 17 Reaction Kinetics Test
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Modern Chemistry Ch 17 Reaction Kinetics Test
 - Fact-Checking eBook Content of Modern Chemistry Ch 17 Reaction Kinetics Test
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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