

AP Chemistry Chapter 14 Answers – Zumdahl

14.27

a.
 $\text{HClO}_4(\text{aq}) + \text{H}_2\text{O}(\text{l}) \rightarrow \text{H}_3\text{O}^+(\text{aq}) + \text{ClO}_4^-(\text{aq})$. Only the forward reaction is indicated since HClO_4 is a strong acid and is basically 100% dissociated in water. For acids, the dissociation reaction is commonly written without water as a reactant. The common abbreviation for this reaction is: $\text{HClO}_4(\text{aq}) \rightarrow \text{H}^+(\text{aq}) + \text{ClO}_4^-(\text{aq})$. This reaction is also called the K_a reaction as the equilibrium constant for the reaction is called K_a .

b.
 Propanoic acid is a weak acid, so it is only partially dissociated in water. The dissociation reaction is: $\text{CH}_3\text{CH}_2\text{CO}_2\text{H}(\text{aq}) + \text{H}_2\text{O}(\text{l}) \rightleftharpoons \text{H}_3\text{O}^+ + \text{CH}_3\text{CH}_2\text{CO}_2^-(\text{aq})$ or $\text{CH}_3\text{CH}_2\text{CO}_2\text{H}(\text{aq}) \rightleftharpoons \text{H}^+(\text{aq}) + \text{CH}_3\text{CH}_2\text{CO}_2^-(\text{aq})$.

c.
 NH_4^+ is a weak acid. Similar to propanoic acid, the dissociation reaction is:



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An acid is a proton (H^+) donor and a base is a proton acceptor. A conjugate acid-base pair differs by only a proton (H^+).

	Acid	Base	Conjugate Base of Acid	Conjugate Acid of Base
a.	H_2CO_3	H_2O	HCO_3^-	H_3O^+
b.	$\text{C}_5\text{H}_5\text{NH}^+$	H_2O	$\text{C}_5\text{H}_5\text{N}$	H_3O^+
c.	$\text{C}_6\text{H}_5\text{NH}^+$	HCO_3^-	$\text{C}_6\text{H}_5\text{N}$	H_2CO_3

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Strong acids have a $K_a \gg 1$ and weak acids have $K_a < 1$. Table 14.2 in the text lists some K_a values for weak acids. K_a values for strong acids are hard to determine so they are not listed in the text. However, there are only a few common strong acids so if you memorize the strong acids, then all other acids will be weak acids. The strong acids to memorize are HCl , HBr , HI , HNO_3 , HClO_4 , and H_2SO_4 .

- HClO_4 is a strong acid.
- HOCl is a weak acid ($K_a = 3.5 \times 10^{-8}$).
- H_2SO_4 is a strong acid.
- H_2SO_3 is a weak diprotic acid with K_{a1} and K_{a2} values less than one.

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The K_a value is directly related to acid strength. As K_a increases, acid strength increases. For water, use K_w when comparing the acid strength of water to other species. The K_a values are:



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SA Adler



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Izutsu,2009-09-22 An excellent resource for all graduate students and researchers using electrochemical techniques After introducing the reader to the fundamentals the book focuses on the latest developments in the techniques and applications in this field This second edition contains new material on environmentally friendly solvents such as room temperature ionic liquids Modern Chemistry Holt Rinehart & Winston,Holt, Rinehart and Winston Staff,2001 *Principles of Modern*

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concentration in the atmosphere of carbon dioxide may be met Cor sion is recognized as having an electrochemical basis The synthesis of nylon now contains an important electrochemical stage Some central biological mechanisms have been shown to take place by means of electrochemical reactions A number of American organizations have recently recommended greatly increased activity in training and research in electrochemistry at universities in the United States Modern University Chemistry, Custom Pub Norbert T. Porile,1993 **Modern Aspects of Electrochemistry** Ralph E. White,Brian E. Conway,Costas G. Vayenas,2012-12-06 This volume of Modern Aspects contains seven chapters The major topics covered in the first six chapters of this volume include fundamentals of solid state electrochemistry kinetics of electrochemical hydrogen entry into metals and alloys oxidation of organics fuel cells electrode kinetics of trace anion catalysis nano structural analysis The last chapter is a corrected version of chapter four from Volume 35 Faisal M Al fageer and Howard W Pickering begin the first chapter by going back to 1864 and Cailletet who found that some hydrogen evolved and was absorbed by iron when it was immersed in dilute sulfuric acid The absorption of hydrogen into metals and alloys can lead to catastrophic failures of structures They discuss the kinetics of electrochemical hydrogen entry into metals and alloys In chapter three Clyde L Briant reviews the electrochemistry corrosion and hydrogen embrittlement of unalloyed titanium He begins by reviewing the basic electrochemistry and general corrosion of titanium He also discusses pitting and galvanostatic corrosion followed by a review of hydrogen embrittlement emphasizing the formation of hydrides and the effect of these on titanium s mechanical properties Christos Comninellis and Gy6rgy F6ti discuss the oxidative electrochemical processes of organics in chapter three They begin by defining direct and indirect electrochemical oxidation of organics They introduce a model that allows them to distinguish between active strong and non active weak anodes Different classes of organic compounds are used for kinetic models of organic oxidation at active and non active type anodes *Modern Aspects of Electrochemistry* Brian E. Conway,Ralph E. White,2006-04-18 Recognized experts present incisive analyses of both fundamental and applied problems in this continuation of a highly acclaimed series Topics in Number 35 include Impedance spectroscopy with specific applications to electrode processes involving hydrogen Fundamentals and contemporary applications of electroless metal deposition The development of computational electrochemistry and its application to electrochemical kinetics Analysis of electrolyte solutions at high concentrations Applications of the Born theory to solvent polarization by ions and its extensions to treatment of kinetics of ionic reactions LIST *Solutions to Problems, Principles of Modern Organic Chemistry* James Cason,1967 **Elementary Modern Chemistry** Wilhelm Ostwald,Harry Wheeler Morse,1909 Student's Guide to Chemistry; a Modern Introduction David Brooks,1974 *American Illustrated Magazine* ,1913 American Magazine ,1913 Solutions to Exercises Roxy Wilson,2002-05 Full solutions to all end of chapter exercises in the text are provided With an instructor s permission this manual may be made available to students **Catechism of Modern Elementary Chemistry Or Solutions of the Questions Set at the London University Matriculation Examinations 1844-82** E. W.

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