

AP Chemistry Chapter 19 Thermodynamics Practice Test

Multiple Choice

Identify the choice that best completes the statement or answers the question.

- _____ 1. ΔG_f° refers to the free-energy change associated with which of the following?
- the formation of 1 mol of a compound from its constituent elements.
 - the decomposition of 1 mole of a compound into its constituent elements.
 - the freezing of 1 mol of a substance.
 - the melting of 1 mol of a substance.
 - all of the above.
- _____ 2. Under what conditions can we absolutely say a system is at equilibrium at constant pressure and temperature?
- $\Delta H = 0$
 - $\Delta H > 0$
 - $\Delta G = 0$
 - $\Delta G > 0$
 - $\Delta S > 0$
- _____ 3. Which of the following substances would you expect to have the highest entropy at 1 atm of pressure?
- $\text{H}_2\text{O}(s)(1 \text{ mol})$
 - $\text{H}_2\text{O}(l)(1 \text{ mol})$
 - $\text{H}_2\text{O}(g)(1 \text{ mol})$
 - 0.5 mol of ethanol in 0.5 mol of $\text{H}_2\text{O}(l)$
 - a gaseous solution of 0.5 mol of $\text{H}_2\text{O}(g)$ and 0.5 mol of $\text{CH}_4(g)$
- _____ 4. For any isolated spontaneous process that is endothermic, which of the following conditions are true
(1) $\Delta G_{\text{system}} < 0$ (2) $\Delta G_{\text{system}} > 0$ (3) $\Delta S_{\text{system}} < 0$ (4) $\Delta S_{\text{system}} > 0$ (5) $\Delta S_{\text{surrounds}} < 0$ (6) $\Delta S_{\text{surrounds}} > 0$
- (1), (3) and (4)
 - (2), (4) and (6)
 - (1), (4) and (5)
 - (1), (4) and (6)
 - (2), (4) and (6)
- _____ 5. For which of the following changes would you expect ΔS to be less than zero?
- $\text{H}_2\text{O}(l) \rightarrow \text{H}_2\text{O}(g)$
 - $\text{CO}(g) + \frac{1}{2}\text{O}_2(g) \rightarrow \text{CO}_2(g)$
 - $2 \text{SO}_3(g) \rightarrow 2 \text{SO}_2(g) + \text{O}_2(g)$
 - $\text{CO}_2(s) \rightarrow \text{CO}_2(g)$
 - $2 \text{AgCl}(s) \rightarrow 2 \text{Ag}(s) + \text{Cl}_2(g)$
- _____ 6. When KNO_3 dissolves in water at room temperature, ΔH is positive for the dissolution process. Given this information, what can you conclude?
- $\Delta G > 0$ for the dissolution process.
 - $\Delta G = 0$ for the dissolution process.
 - The dissolving of salts in water is always a spontaneous process.
 - $\Delta S > 0$ for the dissolution process.
 - $\Delta S < 0$ for the dissolution process.
- _____ 7. Which of the following is true for the sublimation of solid CO_2 at room temperature and pressure?
- ΔH and ΔS are positive.
 - ΔH and ΔS are negative.
 - ΔH is negative and ΔS are positive.
 - ΔG becomes more positive as temperature increases.
 - ΔG is positive.

Modern Chemistry Chapter 19 Work Answers

K Morrison



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